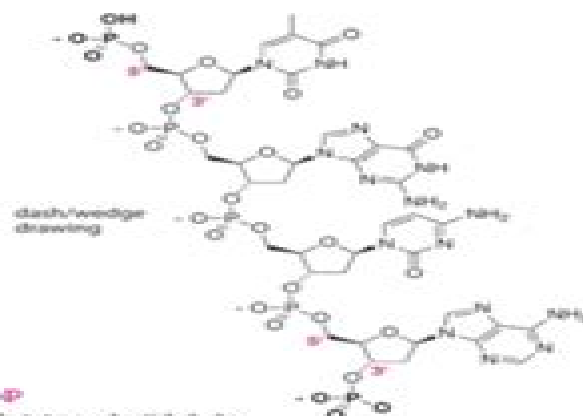


## Answers to Chapter 3 (in-text & asterisked problems)

### Answer 3.1



The labor of drawing a simple tetranucleotide helps one to appreciate why we use a combination of one-letter code and atom-numbering to discuss DNA.

### Answer 3.2

In the gas phase, the charge separation costs 83 kcal mol<sup>-1</sup>, which is much more costly than the maximum benefit of aromaticity (35 kcal mol<sup>-1</sup>). However, in water, where the dielectric constant is about 78, charge separation costs only 1.1 kcal mol<sup>-1</sup>, which is a small compared with the energetic benefits conferred by aromaticity.

$$\begin{aligned}
 \text{Energy} &= k_e \frac{q_1 \cdot q_2}{r \cdot \text{distance}} \\
 &= \frac{9.0 \times 10^9 \cancel{\text{J} \cdot \text{m}} \cdot (1.6 \times 10^{-19} \cancel{\text{C}}) (1.6 \times 10^{-19} \cancel{\text{C}})}{1 \cdot (4.0 \cancel{\text{Å}})} \cdot \frac{1 \text{ kcal}}{1000 \text{ J}} \cdot \frac{1 \text{ m}}{10^{10} \cancel{\text{Å}}} \\
 &= 5.8 \times 10^{-13} \text{ J} \cdot \frac{1 \text{ kcal}}{1000 \text{ J}} \cdot \frac{10^{10}}{\text{mol}} \\
 &= 58 \text{ kJ mol}^{-1} \\
 &\text{or } 83 \text{ kcal mol}^{-1} \quad (1 \text{ cal} = 4.18 \text{ J}) \quad \text{in gas phase } (\epsilon = 1)
 \end{aligned}$$



$$\text{Energy} = 1.1 \text{ kcal mol}^{-1} \quad \text{in water at 25 °C } (\epsilon = 78)$$

It is important to recognize that the positive charge on nitrogen and negative charge on oxygen is a Lewis structure formalism. In fact, the nitrogen atom has a partial negative charge and most of the positive charge is distributed on the various hydrogen atoms. Surprisingly, ab initio calculations HF/STO-3G predict a dipole moment of 4.8 Debye units, not too different from the dipole moment predicted from a +1 charge and -1 charge separate by 4.0 Å (4.0 Debye units).

# Mastering Chemistry Chapter 3 Answer Key

**M Walker**



## **Mastering Chemistry Chapter 3 Answer Key:**

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