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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8 3 mole h₂s and multiply it with 2 mole h₂o over 2 mole h₂s from the equation

$$8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$$

 8.3 mole h₂s x 18g h₂o i mole h₂o 149 4g h₂o is your theoretical yield so percent yield 137 1g h₂o actual yield 149 4g h₂o x 100
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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$

 if the reaction of 21 3 grams of o₂ produces a 47 3 yield how many grams of co would be produced step one figure out how many moles of the limiting reagent you have

$$\frac{21.3 \text{ grams O}_2}{32 \text{ grams mole}} = 0.66 \text{ moles O}_2$$

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 calculated mass of the product expected x100 in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction

$$\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$$

 4 al 3 o₂ 2 al₂o₃ b determine the theoretical yield of al₂o₃ 160 0

g al x 1 mol al x 2 mol al₂O₃ x 101 964 g al₂O₃ 302 3 g al₂O₃ 26 982 g al 4 mol al 1 mol al₂O₃ c determine the percent yield
 yield actual x 100 260 0 g x 100 86 01 theoretical 302 3 g

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 weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12
 c is the limiting reagent 4 33 g of h 2 are left over 13

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 what is the

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