

Chapter 4: Supramolecular Problems - Answers

1. Assign the oxidation number to each atom in the following elements, molecules and ionic substances:

- a) CO_2 b) CH_4 c) HNO_3 d) Ca(OH)_2 e) H_2SO_4 f) Ag
 $\text{C} = +4$ $\text{C} = +4$ $\text{N} = +5$ $\text{Ca} = +2$ $\text{S} = +6$ $\text{Ag} = 0$
 $\text{O} = -2$ $\text{H} = +1$ $\text{O} = -2$ $\text{O} = -2$ $\text{O} = -2$ $\text{Ag} = 0$
- g) Fe h) BaCl_2 i) CO_2 j) H_2O k) HNO_3 l) Fe
 $\text{Fe} = 0$ $\text{Ba} = +2$ $\text{C} = +4$ $\text{H} = +1$ $\text{N} = +5$ $\text{Fe} = 0$
 $\text{Cl} = -1$ $\text{O} = -2$ $\text{O} = -2$ $\text{O} = -2$ $\text{O} = -2$ $\text{Fe} = 0$

2. Assign the oxidation number to each atom in the following ions:

- a) NO^+ b) CO_3^{2-} c) HNO_2 d) SO_4^{2-} e) PO_4^{3-} f) HCO_3^-
 $\text{N} = +2$ $\text{C} = +3$ $\text{N} = +3$ $\text{S} = +6$ $\text{P} = +5$ $\text{C} = +3$
 $\text{O} = -2$ $\text{O} = -2$ $\text{O} = -2$ $\text{O} = -2$ $\text{O} = -2$ $\text{O} = -2$

3. For the following reactions identify if it is an oxidation-reduction reaction. If it is a redox reaction identify the oxidizing agent, the reducing agent, the substance being oxidized and the substance being reduced. Identify the oxidation number change to help to support these assignments.



Redox rxn. Ca oxidized (oxidizing agent); HCl H^+ \longrightarrow H_2
 HCl reduced (red. reducing agent); HCl Cl^- \longrightarrow Cl^-



Not a redox rxn. Ca as Ca^{2+} is both oxidized and reduced



Not a redox rxn. H_2O oxidized (oxid. reducing agent); H_2O as H^+ \longrightarrow H^+
 H_2O reduced (red. reducing agent); H_2O as O^{2-} \longrightarrow O^{2-}



Not a redox reaction. No change in ox. #



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Gerald W. Gibson



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