#### Chemistry: Balancing Chemical Equations

<u>Directions</u>: First, balance each of the chemical equations below. Then, classify each reaction as **synthesis**, **decomposition**, **single-replacement**, or **double-replacement**. To earn full credit, write the words out when classifying.

Balance the equation...

- \_2\_Sb + \_3\_Cl<sub>2</sub> → \_2\_SbCl<sub>3</sub>
- 2. Mg + \_\_\_O; → \_\_2\_MgO
- 3. \_\_\_\_ CaCl<sub>2</sub> → \_\_\_ Ca + \_\_\_ Cl<sub>2</sub>
- 2 NaClO₂ → 2 NaCl + 3 O₂
- Fe + 2 HCl → FeCl<sub>2</sub> + H<sub>2</sub>
- 6. CuO + H₂ → Cu + H₂O
- 7. 2 Al + 3  $H_2SO_4 \rightarrow$  Al<sub>2</sub>( $SO_4$ )<sub>3</sub> + 3  $H_2$
- MgBr<sub>2</sub> + \_\_\_ Cl<sub>2</sub> → \_\_\_ MgCl<sub>2</sub> + \_\_\_ Br<sub>2</sub>
- 9. \_\_\_\_SnO<sub>2</sub> + \_\_2 C → \_\_\_\_Sn + \_\_2 CO
- 10.  $Pb(NO_3)_2 + ___ H_2S \rightarrow ___ PbS + __2 HNO_3$
- 11. \_\_2\_HgO → \_\_2\_Hg + \_\_\_O;
- 12. \_\_2\_ KClO<sub>3</sub> → \_\_2\_ KCl + \_\_3\_ O<sub>2</sub>
- 13.  $N_2 + 3 H_2 \rightarrow 2 NH_3$
- 14. \_\_2 NaBr + \_\_\_ Cl₂ → \_\_2 NaCl + \_\_\_ Br₂
- 15. \_\_\_\_ Zn + \_\_2\_ AgNO<sub>3</sub> → \_\_\_\_ Zn(NO<sub>3</sub>)<sub>2</sub> + \_\_2\_ Ag
- 16. \_\_\_\_ Sn + \_\_2 Cl<sub>2</sub> → \_\_\_\_ SnCl<sub>4</sub>
- 17. \_\_\_\_ Ba(OH)<sub>2</sub> → \_\_\_\_ BaO + \_\_\_\_ H<sub>2</sub>O

...and classify it.

synthesis

synthesis

decomposition

decomposition

single replacement

single replacement

single replacement

single replacement

single replacement

double replacement

decomposition

decompostion

synthesis

single replacement

single replacement

synthesis

decomposition

# **Mastering Chemistry Exercise Answers**

**Thomas Bond, Chris Hughes** 

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