

Name \_\_\_\_\_

## Stoichiometry

21. How do you determine the amount of product formed in a reaction?

*Use stoichiometry based on the limiting reactant.*

22. What is the limiting reactant in  $2SO_2 + O_2 \rightarrow 2SO_3$  if you start with 3 moles of  $SO_2$  and 2 moles of  $O_2$ ?

*$SO_2$*

23. Why is it important to balance a chemical equation for stoichiometry calculations?

*To ensure the correct mole ratios are used.*

24. How can you increase the percent yield of a reaction?

*Optimize reaction conditions and minimize losses.*

25. What information is needed to calculate theoretical yield?

*Balanced chemical equation and amounts of reactants.*

26. If the percent yield is 50% and the actual yield is 5 grams, what is the theoretical yield?

*10 grams*

27. What is the first step in stoichiometric calculations?

*Balance the chemical equation.*

28. In  $C_3H_8 + 5O_2 \rightarrow 3CO_2 + 4H_2O$ , what is the mole ratio of  $O_2$  to  $CO_2$ ?

*5:3*

29. If you have 10 grams of  $NaCl$  and the reaction has a percent yield of 75%, how much product is obtained?

*7.5 grams*

30. How do you identify the limiting reactant in a multi-step reaction?

*Determine the limiting reactant for each step and use it for calculations in subsequent steps.*



# Holt Chemistry Study Guide Concept Review

## Stoichiometry

**John W. Moore, Conrad L.  
Stanitski, Peter C. Jurs**



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