

## Limiting Reagents and Percentage Yield Worksheet

### 1. Consider the reaction



- a) 80.0 grams of iodine(V) oxide,  $\text{I}_2\text{O}_5$ , reacts with 28.0 grams of carbon monoxide,  $\text{CO}$ . Determine the mass of iodine  $\text{I}_2$ , which could be produced?

80 g $\text{I}_2\text{O}_5$	1 mol $\text{I}_2\text{O}_5$	1 mol $\text{I}_2$		XS
1	333.8 g $\text{I}_2\text{O}_5$	1 mol $\text{I}_2\text{O}_5$		
28 g $\text{CO}$	1 mol $\text{CO}$	1 mol $\text{I}_2$	253.8 g $\text{I}_2$	
1	28 g $\text{CO}$	5 mol $\text{CO}$	1 mol $\text{I}_2$	= 50.8 g $\text{I}_2$

- b) If, in the above situation, only 0.160 moles, of iodine,  $\text{I}_2$  was produced.  
i) what mass of iodine was produced?

0.160 mol $\text{I}_2$	253.8 g $\text{I}_2$	= 40.6 g $\text{I}_2$
	1 mol $\text{I}_2$	

- ii) what percentage yield of iodine was produced.

40.6 g $\text{I}_2$	$\times 100 =$	79.9 %
50.8 g $\text{I}_2$		

### 2. Zinc and sulphur react to form zinc sulphide according to the equation.



If 25.0 g of zinc and 30.0 g of sulphur are mixed,

- a) Which chemical is the limiting reactant?

25.0 g Zn	1 mol Zn	1 mol ZnS	= 382 mol ZnS	LR
1	65.41 g Zn	1 mol Zn		
30.0 g S	1 mol S	1 mol ZnS	= 934 mol ZnS	XS
1	32.1 g S	1 mol S		

- b) How many grams of  $\text{ZnS}$  will be formed?

382 mol ZnS	97.51 g ZnS	= 37.2 g ZnS
	1 mol ZnS	

- c) How many grams of the excess reactant will remain after the reaction is over?

$$.934 \text{ mol ZnS} - .382 \text{ mol ZnS} = .552 \text{ mol ZnS}$$

.552 mol ZnS	1 mol S	32.1 g S	= 17.7 g S
	1 mol ZnS	1 mol S	

### 3.

# Limiting Reactant And Percent Yield Answers

**Padhraic Smyth**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

## **Limiting Reactant And Percent Yield Answers:**

## Decoding **Limiting Reactant And Percent Yield Answers**: Revealing the Captivating Potential of Verbal Expression

In a period characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its ability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Limiting Reactant And Percent Yield Answers**," a mesmerizing literary creation penned by a celebrated wordsmith, readers embark on an enlightening odyssey, unraveling the intricate significance of language and its enduring effect on our lives. In this appraisal, we shall explore the book's central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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