

Limiting Reagent & Percent Yield Practice Worksheet

1. When copper (II) chloride reacts with sodium nitrate, copper (II) nitrate and sodium chloride are formed.

a. Write the balanced equation for the reaction given above:



- b. If 15 grams of copper (II) chloride react with 20 grams of sodium nitrate, how much sodium chloride can be formed?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.362 \text{ g CuCl}_2} \times \frac{2 \text{ mol NaCl}}{1 \text{ mol CuCl}_2} \times \frac{58.443 \text{ g NaCl}}{1 \text{ mol NaCl}} = 13 \text{ g NaCl}$$

$$20 \text{ g NaNO}_3 \times \frac{1 \text{ mol NaNO}_3}{84.997 \text{ g NaNO}_3} \times \frac{2 \text{ mol NaCl}}{2 \text{ mol NaNO}_3} \times \frac{58.443 \text{ g NaCl}}{1 \text{ mol NaCl}} = 13.8 \text{ g NaCl}$$

Only 13 g can be formed

- c. What is the limiting reagent for the reaction in #2? copper (II) chloride

- d. How many grams of copper (II) nitrate is formed?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.452 \text{ g CuCl}_2} \times \frac{1 \text{ mol Cu(NO}_3)_2}{1 \text{ mol CuCl}_2} \times \frac{187.467 \text{ g Cu(NO}_3)_2}{1 \text{ mol Cu(NO}_3)_2} = 20.9 \text{ g Cu(NO}_3)_2$$

- e. How much of the excess reagent is left over in this reaction?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.452 \text{ g CuCl}_2} \times \frac{2 \text{ mol NaNO}_3}{1 \text{ mol CuCl}_2} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 19 \text{ g NaNO}_3$$

$$20 \text{ g NaNO}_3 - 19 \text{ g NaNO}_3 = 1 \text{ g NaNO}_3$$

- f. If 11.3 grams of sodium chloride are formed in the reaction described in problem #2, what is the percent yield of this reaction?

$$\text{Percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100 \text{ percent}$$

$$11.3/13.0 \times 100\% = 86.9\%$$

2. When lead (II) nitrate reacts with sodium iodide, sodium nitrate and lead (II) iodide are formed.

a. Balance the following equation:



- b. If I start with 25.0 grams of lead (II) nitrate and 15.0 grams of sodium iodide, how many grams of sodium nitrate can be formed?

$$25.0 \text{ g Pb(NO}_3)_2 \times \frac{1 \text{ mol Pb(NO}_3)_2}{331.214 \text{ g Pb(NO}_3)_2} \times \frac{2 \text{ mol NaNO}_3}{1 \text{ mol Pb(NO}_3)_2} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 12.8 \text{ g NaNO}_3$$

$$15.0 \text{ g NaI} \times \frac{1 \text{ mol NaI}}{149.89 \text{ g NaI}} \times \frac{2 \text{ mol NaNO}_3}{2 \text{ mol NaI}} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 8.51 \text{ g NaNO}_3$$

8.51 g NaNO₃ can be made

Limiting Reagent And Percent Yield Answers

Mike Jess



Limiting Reagent And Percent Yield Answers:

Calculating Theoretical Yield Is Not Impossible! Michael S. Leonard, 2025-02-25 Calculating theoretical and percent yield is a fundamental skill for the laboratory. This book primarily targets Organic Chemistry Laboratory courses at the high school or college and university level as a supplemental resource to help students master this skill. It begins with simple examples from everyday life, demonstrates the importance of balancing the equation, addresses the role of the mole in these computations, discusses different types of liquids, considers the role of significant figures, and culminates with the planning of syntheses. There are suggestions for further reading as well as practice problems and questions to ensure mastery. Begins with examples from everyday life that enable students to understand the concepts of theoretical and percent yield before applying those concepts to the laboratory. Addresses the necessity of balancing the reaction equation, the centrality of the mole in these calculations, and the role of significant figures in reporting the answer. Explains how to approach the calculations when using neat liquids or solutions. The culmination of this text is the use of the same thought processes to plan the amounts of reactants needed for syntheses of desired quantities of product. All of the problems in the book include detailed solutions with accompanying text to explain the answers, and ancillaries also include suggestions for further reading.

Chemistry: The Central Science Theodore L. Brown, H. Eugene LeMay Jr., Bruce E. Bursten, Catherine Murphy, Patrick Woodward, Steven Langford, Dalius Sagatys, Adrian George, 2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text, think again. In response to market request, we have created the third Australian edition of the US bestseller Chemistry: The Central Science. An extensive revision has taken this text to new heights. Triple checked for scientific accuracy and consistency, this edition is a more seamless and cohesive product yet retains the clarity, innovative pedagogy, functional problem solving, and visuals of the previous version. All artwork and images are now consistent in quality across the entire text. And with a more traditional and logical organisation of the Organic Chemistry content, this comprehensive text is the source of all the information and practice problems students are likely to need for conceptual understanding, development of problem solving skills, reference, and test preparation.

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Access to Chemistry Avril Higton, Mike Clemmet, Elaine Golding, Alan V Jones, 2007-10-31 This text has been specifically designed to prepare people with previously limited chemical knowledge for entrance into science related courses such as Foundation and Access courses which involve chemistry in higher education Until now there have been no texts available for use on these courses and this book fills that gap Access to Chemistry effectively forms a self study course which is split into separate modules and units covering the full spread of concepts required for those needing a basic knowledge of chemistry The material is presented in a friendly and easy to use manner which allows the student to pace their acquisition of knowledge and gain increasing confidence in order to succeed in understanding essential relevant concepts Other useful features of this book include starter diagnostic tests worked examples and self study tests with answers at the end of each unit In addition to Access or Foundation course students and their tutors to whom this book will prove essential it will have an appeal also as a revision text for those needing a refresher after a break in the subject In addition it will be of interest to members of the general public who wish to better educate themselves on chemical matters as it provides a clear and useful insight into areas such as health home chemicals business market trends and gardening

Concepts And Problems In Physical Chemistry P.S. Raghavan, 1997 Contents Introduction Atoms Molecules and Formulas Chemical Equations and Stoichiometry Aqueous Reactions and Solution Stoichiometry Gases Intermolecular Forces Liquids and Solids Atoms Structure and the Periodic Table Chemical Bonding Chemical Thermodynamics Solutions Chemical Kinetics Chemical Equilibrium Acids and Bases Ionic Equilibria I Ionic Equilibria II Redox Reactions Electrochemistry Nuclear Chemistry

Meta-Analytics Steven Simske, 2019-03-10 Meta Analytics Consensus Approaches and System Patterns for Data Analysis presents an exhaustive set of patterns for data science to use on any machine learning based data analysis task The book virtually ensures that at least one pattern will lead to better overall system behavior than the use of traditional analytics approaches The book is meta to analytics covering general analytics in sufficient detail for readers to engage with and understand hybrid or meta approaches The book has relevance to machine translation robotics biological and social sciences medical and healthcare informatics economics business and finance In addition the analytics within can be applied to predictive algorithms for everyone from police departments to sports analysts Provides comprehensive and systematic coverage of machine learning based data analysis tasks Enables rapid progress towards competency in data analysis techniques Gives exhaustive and widely applicable patterns for use by data scientists Covers hybrid or meta approaches along with general analytics Lays out information and

practical guidance on data analysis for practitioners working across all sectors

Principles of Modern Chemistry David W. Oxtoby, H. Pat Gillis, Laurie J. Butler, 2016-01-01 Long considered the standard for honors and high level mainstream general chemistry courses PRINCIPLES OF MODERN CHEMISTRY continues to set the standard as the most modern rigorous and chemically and mathematically accurate text on the market This authoritative text features an atoms first approach and thoroughly revised chapters on Quantum Mechanics and Molecular Structure Chapter 6 Electrochemistry Chapter 17 and Molecular Spectroscopy and Photochemistry Chapter 20 In addition the text utilizes mathematically accurate and artistic atomic and molecular orbital art and is student friendly without compromising its rigor End of chapter study aids focus on only the most important key objectives equations and concepts making it easier for students to locate chapter content while applications to a wide range of disciplines such as biology chemical engineering biochemistry and medicine deepen students understanding of the relevance of chemistry beyond the classroom

Physical Chemistry for JEE Advanced: Part 1, 3E (Free Sample) K. S. Verma, 2022-05-19 Physical Chemistry for JEE Advanced Part 1 a Cengage Exam Crack Series product is designed to help aspiring engineers focus on the subject of physical chemistry from two standpoints To develop their caliber aptitude and attitude for the engineering field and profession To strengthen their grasp and understanding of the concepts of the subjects of study and their applicability at the grassroots level Each book in this series approaches the subject in a very conceptual and coherent manner While its illustrative solved examples facilitate easy mastering of the concepts and their applications an array of solved problems exposes the students to a variety of questions that they can expect in the examination The coverage and features of this series of books make it highly useful for all those preparing for JEE Main and Advanced and aspiring to become engineers

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Investigating Chemistry Matthew Johll, 2008-12-22 In its new second edition *Investigating Chemistry A Forensic Science Perspective* remains the only book that uses the inherently fascinating topics of crime and criminal investigations as a context for teaching the fundamental chemical concepts most often covered in an introductory nonmajors course. Covering all the standard topics, Matthew Johll capitalizes on the surge of interest in the scientific investigation of crime as sparked by CSI and other television shows, bringing together the theme of forensic science and the fundamentals of chemistry in ways that are effective and accessible for students. This edition features refined explanations of the chemical concepts which are the core of the book, as well as a more thoroughly integrated forensic theme, updated features, and an expanded media supplements package.

CHEMISTRY HANDBOOK & STUDY GUIDE Gr11-12 NE Kevin Smith, 2024-02-01 A comprehensive summary of Grade 11/12 Physics. Simple logical summaries with example exam questions and work through solutions. The book covers the fundamentals of Grade 11/12 Physics and complements the material in any class text.

Chemistry Rob Lewis, Wynne Evans, 2018-03-31 The fifth edition of this engaging and established textbook provides students with a complete course in chemical literacy and assumes minimal prior experience of science and maths. Written in an accessible and succinct style, this book offers comprehensive coverage of all the core topics in organic, inorganic, and physical chemistry. Topics covered include bonding, moles, solutions, and solubility, energy changes, equilibrium, organic compounds, and spectroscopy. Each unit contains in-text exercises and revision questions to consolidate learning at every step and is richly illustrated with diagrams and images to aid understanding. This popular text is an essential resource for students who are looking for an accessible introductory textbook. It is also ideal for non-specialists on courses such as general science, engineering, environmental health, or life sciences. New to this Edition: A foreword by Professor Sir John Meurig Thomas FRS, former Director of the Royal Institution. Three additional units on Gibbs Energy, Changes, Organic Mechanisms, and Fire and Flame. Accompanying online resources for this title can be found at bloomsburyonlineresources.com/chemistry/5e. These resources are designed to support teaching and learning when using this textbook and are available at no extra cost.

General Chemistry Ralph H. Petrucci, William S. Harwood, 1993 *General Chemistry Principles and Modern Applications* is recognized for its superior problems, lucid writing, and precision of argument. This updated and expanded edition retains the popular and innovative features of previous editions, including Feature Problems, Follow-up Integrative and Practice Exercises to accompany every chapter, Example and Focus On application boxes, as well as new Keep in Mind marginal notes. Topics covered include atoms and the atomic theory, chemical compounds and reactions, gases, Thermochemistry, electrons in atoms, chemical bonding, liquids, solids, and intermolecular forces, chemical kinetics, principles of chemical equilibrium, acids and bases, electrochemistry, representative and transitional elements, and nuclear and organic chemistry. For individuals interested in a broad overview of chemical principles and applications.

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