Worksheet Specific Heat

Specific Heat of water =4.18 J/g °C

- A sample of mercury is heated from 25.5 °C to 52.5 °C. In the process, 3050J of heat are absorbed.
 What mass of mercury was in the sample? The specific heat of mercury is 0. 140 J/g°C. ANS: 807 g.
- A block of aluminum with a mass of 140g is cooled from 98.4°C to 62.2°C with a release of 1137J of heat. From these data, calculate the specific heat of aluminum. ANS: 0.224
- A total of 226 J of heat are absorbed as 58.3 g of lead is heated from 12.0°C to 42.0°C. From this
 data, what is the specific heat of lead? ANS: 0.129
- What is the change in heat energy when 132g of copper at 32.2°C is raised to 45.0°C? The specific heat of copper is 0.385 J/g °C. ANS: 650 J
- What is the change in heat energy when 40.0g of water at 25.0°C is raised to 60.0°C? ANS: 5852 J.
- What is the change in heat energy when 75.0 of water at 9.8°C is raised to 22.4°C? 3950 J
- What is the change in heat energy when 125g of water at 10.0°C is raised to 90. 0°C? ANS: 41.8 kJ
- What is the change in heat energy when 64.82 g of aluminum metal at 100.0°C is cooled to 82.0 °C?
 The specific heat of aluminum is 0.897 J/g °C. ANS: 1047
- What is the change in temperature if 24g of water at 50°C absorbs 985 J of heat energy? ANS: 9.82
- 10. What is the mass of a sample of iron if that sample lost 2300J of heat energy when it cooled from 80 °C to 30°C? The specific heat of iron is 0.449 J/g °C. ANS: 102 g
- What is the specific heat of metal if its mass is 26.86g and it requires 418.6 J of heat energy to raise its temperature from 27.4 °C to 67.3°C? ANS: 0.391
- How much heat energy is needed to raise the temperature of 59.7g of cadmium from 25°C to 100°C? The specific heat of cadmium is 0.231 J/g °C. ANS: 1034 J

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N. Hawas

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