

Le Chatelier's Principle Worksheet

1) For the reaction below, which way would the equilibrium shift, for each situation below, to the right or to the left?



- (a) Decrease the concentration of dihydrogen sulfide. moves to the left
- (b) Increase the pressure on the system. Moves to Left
- (c) Increase the temperature of the system. Moves to Right
- (d) Increase the concentration of carbon disulfide. Moves to right
- (e) Decrease the concentration of methane (CH_4). Moves to left

2) What would happen to the position of the equilibrium when the following changes are made to the equilibrium system below?



- (a) Sulfur dioxide is added to the system. The reactants are favored, thus concentration increases.
- (b) Sulfur trioxide is removed from the system. The reactants are favored, thus concentration increases.
- (c) Oxygen is added to the system. The reactants are favored, thus concentration increases.

3) What would happen to the position of the equilibrium when the following changes are made to the reaction below?



- (a) HgO is added to the system. Here the products are favored and the forward reaction is strong.
- (b) The pressure on the system increases. Here the products are favored, but the reverse reaction is stronger.

Le Chatliers Principle Answers Chemfiesta

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