#### AP Biology Study Guide Unit 1

#### Properties of Water:

- 1. Polarity allows H20 to form Hydrogen Bonds
- 2. Cohesion Water H bonds to itself, creating surface tension
- 3. Adhesion Water adheres to other polar substances
  - a. Allows easy transport of water against Gravity
- 4. Universal Solvent Water can dissolve Polar and Ionic Compounds
  - a. This allows for reactions between Aqueous Solutions
- 5. Expands when Frozen this means that Ice floats
- High Specific Heat water can absorb or release a large amount of heat with only a slight change in its own temperature (caused by cohesion)
  - High Heat of Vaporization only warmest molecules can evaporate, so this causes evaporative cooling
- 7. Can be regulated by non-polar hydrophobic substances
- Neutral PH 
   — can form acids or bases if substances are added

Polarity - Difference in electronegativity between atoms bonded together create slightly charged ends of the molecules they form

Hydrophobic Substance - Nonpolar substance that repels water Hydrophilic Substance - Polar substance that attracts water molecules

#### CHON - Carbon, Hydrogen, Nitrogen, Oxygen

- Makeup 99% of living matter
- Strong Biological Cornerstones Because:
  - Smallest Diameter/Valency for atoms of 1, 2, 3, 4 valence electrons
  - Form Stable Covalent Bonds
  - Can form double and triple bonds
  - Versatile

#### Biological Transition Metals - Mn. Fe. Co. Cu. Zn.

Can exist in multiple oxidation states → can be both oxidized and reduced

#### Biological Ions: Na+, Mg2+, Cl-, K+, Ka2+

Used to maintain homeostasis and osmotic concentration

Phosphorous/Sulphur - Energy Carriers which form strong high energy bonds

#### 5 Foundation Biological Molecules:

# **Guide Line Study For Ap Biology Campbell**

Antonia Vlahou, Fulvio Magni, Harald Mischak, Jerome Zoidakis

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