

Temperature, Energy, and Heat Study Guide

I. Basic Definitions

- **Temperature:**
 - A measure of the average kinetic energy of the particles in a substance.
 - Indicates how hot or cold something is.
 - Measured in Celsius ($^{\circ}\text{C}$), Fahrenheit ($^{\circ}\text{F}$), or Kelvin (K).
 - SI unit is Kelvin (K).
- **Heat:**
 - The transfer of thermal energy between objects at different temperatures.
 - Flows from hotter to colder objects.
 - Measured in joules (J).
- **Thermal Energy:**
 - The total kinetic and potential energy of all the particles in a substance.
 - Depends on temperature and the amount of substance.
- **Absolute Zero:**
 - The theoretical temperature at which all molecular motion stops (0 K, -273.15°C).

II. Temperature Scales

- **Celsius ($^{\circ}\text{C}$):**
 - Water freezes at 0°C and boils at 100°C .
- **Fahrenheit ($^{\circ}\text{F}$):**
 - Water freezes at 32°F and boils at 212°F .
- **Kelvin (K):**
 - Absolute temperature scale.
 - 0 K is absolute zero.[†]
 - $\text{K} = ^{\circ}\text{C} + 273.15$

III. Heat Transfer

- **Conduction:**
 - Transfer of heat through direct contact.
 - Occurs primarily in solids.
 - Good conductors (metals) transfer heat quickly.
 - Insulators (wood, plastic) resist heat transfer.
- **Convection:**
 - Transfer of heat through the movement of fluids (liquids or gases).
 - Hotter, less dense fluids rise; cooler, denser fluids sink.
 - Creates convection currents.

Heat Study Guide

Frank P. Incropera, David P. DeWitt



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Transfer 5th Edition provides a complete introduction to the physical origins of heat and mass transfer. Noted for its crystal clear presentation and easy to follow problem solving methodology, Incropera and Dewitt's systematic approach to the first law develops reader confidence in using this essential tool for thermal analysis. **Air Force Manual** United States. Department of the Air Force, 1969

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