

# Hydrolysis -

## The Reaction of Ions with Water

### Lab Report

#### Purpose

The purpose of this lab is to predict whether a substance is acidic, basic, or neutral and to use the universal indicator solution to test it.

#### Hypothesis

Sodium Acetate	$\text{NaCH}_3\text{COO}_{(s)} \rightarrow \text{Na}^+_{(aq)} + \text{CH}_3\text{COO}^-_{(aq)}$ $\text{CH}_3\text{COO}^-_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{CH}_3\text{COOH}_{(aq)} + \text{OH}^-_{(aq)}$		Basic
Sodium Chloride	$\text{NaCl}_{(s)} \leftrightarrow \text{Na}^+_{(aq)} + \text{Cl}^-_{(aq)}$		Neutral
Ammonium Chloride	$\text{NH}_4\text{Cl}_{(s)} \rightarrow \text{NH}_4^+_{(aq)} + \text{Cl}^-_{(aq)}$ $\text{NH}_4^+_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{NH}_4(\text{OH})_{(aq)} + \text{H}^+_{(aq)}$		Acidic
Ammonium Sulfate	$(\text{NH}_4)_2\text{SO}_{4(s)} \rightarrow \text{NH}_4^+_{(aq)} + \text{SO}_4^{2-}_{(aq)}$ $\text{NH}_4^+_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{NH}_4(\text{OH})_{(aq)} + \text{H}^+_{(aq)}$ $\text{SO}_4^{2-}_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{HSO}_4^-_{(aq)} + \text{OH}^-_{(aq)}$	$K_a = 5.4 \times 10^{-10}$	Acidic
		$K_b = 8.3 \times 10^{-13}$	
Aluminum Chloride	$\text{AlCl}_{3(s)} \rightarrow \text{Al}^{3+}_{(aq)} + \text{Cl}^-_{(aq)}$ $\text{Al}^{3+}_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{Al}(\text{OH})_{(aq)} + \text{H}^+_{(aq)}$		Acidic
Calcium Nitrate	$\text{Ca}(\text{NO}_3)_{2(s)} \rightarrow \text{Ca}^{2+}_{(aq)} + \text{NO}_{3(aq)}$		Neutral
Iron(III) Sulfate	$\text{Fe}_2(\text{SO}_4)_{3(s)} \rightarrow \text{Fe}^{3+}_{(aq)} + \text{SO}_4^{2-}_{(aq)}$ $\text{Fe}^{3+}_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{Fe}(\text{OH})^{2+}_{(aq)} + \text{H}^+_{(aq)}$ $\text{Fe}^{3+}_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{HFe}^{2+}_{(aq)} + \text{OH}^-_{(aq)}$ $\text{SO}_4^{2-}_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{HSO}_4^-_{(aq)} + \text{OH}^-_{(aq)}$	$K_a = 6.0 \times 10^{-3}$	Acidic
		$K_b = 1.7 \times 10^{-12}$	
		$K_b = 8.3 \times 10^{-13}$	
Sodium Carbonate	$\text{Na}_2\text{CO}_{3(s)} \rightarrow \text{Na}^+_{(aq)} + \text{CO}_3^{2-}_{(aq)}$ $\text{CO}_3^{2-}_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{HCO}_3^-_{(aq)} + \text{OH}^-_{(aq)}$		Basic

# Heath Chemistry Lab Experiments Answers

**LL Leslie**



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