Step 1 2000 - NoOn

Step 2 N<sub>6</sub>O<sub>6</sub> + H<sub>2</sub> - N<sub>2</sub>O + H<sub>2</sub>O

Step 3  $N_2O + H_2 \rightarrow N_2 + H_2O$ 

Which of the following substances is a reaction intermediate?

A) N<sub>2</sub>O<sub>6</sub> B) NO C) H<sub>2</sub>O D) N<sub>2</sub>O E) N<sub>2</sub>O<sub>6</sub> and N<sub>2</sub>O

15. The value of AH\* for the reaction below is -126 kj. \_\_\_\_kj is released when 2.00 mol of NuOH is formed in the reaction?

2Na<sub>2</sub>O<sub>2m</sub> + 2H<sub>2</sub>O<sub>3m</sub> - 4NaOH<sub>3m</sub> + O<sub>2m</sub>

A) 252 B) 63 C) 3/9 D) 7/8

16. Estimate the enthalpy change for the reaction below from the average bond energies given. There are two G—GI and two G—H bonds in GH<sub>2</sub>G<sub>6</sub>. Remember that energy is absorbed when bonds are broken and released when they are formed.

CH4(g) + 2Cl4(g) -+ CH4(Cl4(g) + 2H4Cl(g)

Average Bond Energies

G-H 41.3 kt/mol Cl-Cl 242 kt/mol

H-C1 432 kt/mat C-Cl 339 kt/mat

A) -232 k[/mol B) +578 k[/mol C) +232 k]/mol

D) -578 kl/mol H) +541 kl/mol

According to the reaction profile below, what is ΔH for the reactions
 H<sub>2</sub>O (q) + 2 Be<sub>b</sub> (q) → 6 HBr (g) + O<sub>b</sub> (g)?



A)-434 M B) 434 M C) -276 M D) 276 M E) 158 M

18. Given the following at AFP (25°C and 1.00 atm)-

 $1/2N_{\rm s}(g) + O_{\rm s}(g) \rightarrow NO_{\rm s}(g) 33.2 \text{ M}$ 

My(g) + 20y(g) - MyOy(g) 11.1 M

Calculate the AH\* for the reaction below at 25°C

2 NO. (60) -+ N. Cute)

A) +11.0 k(fl) +44.3 k(c) +55.3 k(fl) -22.1 k(fl) -55.3 k(

19. Consider the following PE diagram for a catalyzed and uncatalyzed reaction:

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**Daniel F McAuley** 

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