

Chapter 12.1: What is stoichiometry?

Chemical reactions stop when one of the reactants is used up.

Stoichiometry: the study of quantitative relationships between amounts of reactants used and products formed by a chemical reaction.

- Based on the law of conservation of mass
- Mass of reactants must equal the mass of the products
- A balanced equation shows the relationship between the amount of reactant and the amount of product
- We can use a balanced equation to determine the masses involved in a chemical reaction



What we know from the equation:

- 4 molecules of iron reacting with 3 molecules of oxygen to produce 2 molecules of iron (III) oxide
- 4 moles of iron reacting with 3 moles of oxygen to produce 2 moles of iron (III) oxide
- To determine the mass of the substances, multiply by molar mass.

$$4 \text{ mol} \times \frac{55.9 \text{ g Fe}}{1 \text{ mol}} = 223.6 \text{ g Fe}$$

$$3 \text{ mol} \times \frac{32 \text{ g O}_2}{1 \text{ mol}} = 96 \text{ g O}_2$$

$$\text{Total mass of reactants} = 319.6 \text{ g}$$

$$2 \text{ mol} \times \frac{159.8 \text{ g Fe}_2\text{O}_3}{1 \text{ mol}} = 319.6 \text{ g}$$

Why mole ratios are important:

Mole ratios are important in calculations based upon a chemical equation.

- If you know the amount of any one substance, you can calculate the amounts for all the others.

Practice problem

Combustion of propane (C₃H₈)



Calculate the mass of each reactant and each product. Show that the law of conservation of mass is true.

Mole ratio

Mole ratio is the ratio between the numbers of moles of any two substances in a balanced chemical equation.

Example:



Write the mole ratios for each term in the equation.

$$\frac{2 \text{ mol Al}}{3 \text{ mol Br}_2} \qquad \frac{2 \text{ mol Al}}{2 \text{ mol AlBr}_3}$$

$$\frac{3 \text{ mol Br}_2}{2 \text{ mol Al}} \qquad \frac{3 \text{ mol Br}_2}{2 \text{ mol AlBr}_3}$$

$$\frac{2 \text{ mol AlBr}_3}{2 \text{ mol Al}} \qquad \frac{2 \text{ mol AlBr}_3}{3 \text{ mol Br}_2}$$

To find out how many mole ratios there are, multiply the number of terms by the next lower number.

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