

Matric Final Exam Time Table 2025

South Africa Page 1



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Basic Education
REPUBLIC OF SOUTH AFRICA



2025 NATIONAL SENIOR CERTIFICATE (NSC) OCTOBER/NOVEMBER EXAMINATIONS TIMETABLE

WEEK 1	09:00	14:00
Monday 20/10	NON-EXAMINATION DAY	
Tuesday 21/10	Computer Applications Tech P1 (3hrs) Practical	Hindi, Gujarati, Tamil, Telegu, Urdu HL P1 (2hrs) Hindi, Gujarati, Tamil, Telegu, Urdu FAL P1 (2hrs) Hindi, Gujarati, Tamil, Telegu, Urdu SAL P1 (2hrs) Hebrew SAL P1 (2hrs) German HL, SAL P1 (2hrs)
Wednesday 22/10	Information Technology P1 (3hrs) Practical	Arabic, French, Italian, Mandarin, Modern Greek, Serbian, Spanish SAL P1 (2hrs) Latin SAL P1 (3hrs) Portuguese HL, FAL, SAL P1 (2hrs)
Thursday 23/10	English HL P3 (3hrs) English FAL P3 (2½hrs) English SAL P3 (2½hrs)	Arabic, French, Italian, Mandarin, Modern Greek, Serbian, Spanish SAL P2 (2hrs) Latin SAL P2 (2hrs) Portuguese HL P2 (2½hrs), FAL, SAL P2 (2hr) Equine Studies (3hrs)
Friday 24/10	Afrikaans HL P3 (3hrs) Afrikaans FAL P3 (2½hrs) Afrikaans SAL P3 (2½hrs)	Hindi, Gujarati, Tamil, Telegu, Urdu HL P2 (2½hrs) Hindi, Gujarati, Tamil, Telegu, Urdu FAL P2 (2hrs) Hindi, Gujarati, Tamil, Telegu, Urdu SAL P2 (2hrs) Hebrew SAL P2 (2hrs) German HL P2 (2½hrs), SAL P2 (2hrs)
WEEK 2	09:00	14:00
Monday 27/10	isiZulu, isiXhosa, siSwati, isiNdebele HL P3 (3hrs), FAL P3 (2½hrs), SAL P3 (2½hrs)	Accounting P1 (2hrs)
Tuesday 28/10	Sepedi, Sesotho, Setswana, Xitsonga, Tshivenda HL P3 (3hrs), FAL P3 (2½hrs), SAL P3 (2½hrs) South African Sign Language HL P3 (3hrs)	Civil Technology (3hrs) Sport and Exercise Science (3hrs)
Wednesday 29/10	English HL P1 (2hrs) English FAL P1 (2hrs) English SAL P1 (2hrs)	Hindi, Gujarati, Tamil, Telegu, Urdu HL P3 (2½hrs) FAL P3 (2½hrs) Portuguese, German HL P3 (2½hrs) Portuguese FAL P3 (2½hrs)
Thursday 30/10	History P1 (3hrs)	Engineering Graphics and Design P1 (3hrs)
Friday 31/10	Mathematics P1 (3hrs) Mathematical Literacy P1 (3hrs) Technical Mathematics P1 (3hrs)	Religion Studies P1 (2hrs)
WEEK 3	09:00	14:00
Monday 03/11	Mathematics P2 (3hrs) Mathematical Literacy P2 (3hrs) Technical Mathematics P2 (3hrs)	Agricultural Sciences P1 (2½hrs)
Tuesday 04/11	isiZulu, isiXhosa, siSwati, isiNdebele HL P1 (2hrs), FAL P1 (2hrs), SAL P1 (2hrs)	Accounting P2 (2hrs) Maritime Economics (3hrs)
Wednesday 05/11	Sepedi, Sesotho, Setswana, Xitsonga, Tshivenda HL P1 (2hrs), FAL P1 (2hrs), SAL P1 (2hrs) South African Sign Language HL P1 (2hrs)	Religion Studies P2 (2hrs) Mechanical Technology (3hrs)
Thursday 06/11	isiZulu, isiXhosa, siSwati, isiNdebele HL P2 (2½hrs), FAL P2 (2½hrs), SAL P2 (1½hrs)	Business Studies P1 (2hrs)
Friday 07/11	Physical Sciences (Physics) P1 (3hrs) Technical Sciences P1 (3hrs)	History P2 (3hrs)



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web introduction this laboratory exercise relies on a titration technique to determine an unknown concentration of monoprotic acid in solution in the process of titration a basic solution is gradually added to the acidic solution until complete neutralization is obtained

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web with a strong acid and strong base titration strong acids and strong bases refer to species that completely dissociate to form ions in solution by contrast weak acids and bases ionize only partially and the ionization reaction is reversible

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web aug 30 2022 acid base titrations are usually used to find the amount of a known acidic or basic substance through acid base reactions the analyte titrand is the solution with an unknown molarity the reagent titrant is the solution with a known molarity that will react with the analyte

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web conclusion the purpose of the experiment was to identify the unknown solid acid 17 by calculating its molar mass this was done by titration of the acid with sodium hydroxide in order to find the volume of sodium hydroxide required to reach the end point and that value was used in unit analysis to find the molar mass and identify the unknown solid acid

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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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web goals understand that parts of an acid base titration be able to determine the K_a or K_b from pH data associated with the

titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate K_{a1} and K_{a2} for a polyprotic acid by the end of this lab students should be able to

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web titration experiment in the neutralization of hydrochloric acid by sodium hydroxide the mole ratio of acid to base is 1 1
 $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$ 1 mole of HCl would be fully neutralized by one mole of NaOH

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web conclusion the identification of an unknown solid acid the purpose of this experiment was to determine the standardization of sodium hydroxide NaOH and oxalic acid $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ and the molar mass of an unknown solid acid using the molar mass of the standardized sodium hydroxide solution

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web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown acid or base is strong or weak pK_a of an unknown acid or pK_b of the unknown base key terms titration a process where a solution of known strength is added to a certain volume of a treated sample containing an

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being NaOH reaction with a 2nd substance $\text{H}_2\text{C}_2\text{O}_4 + 2\text{NaOH} \rightarrow \text{Na}_2\text{C}_2\text{O}_4 + 2\text{H}_2\text{O}$ to differentiate an acid from a base in a solution a pH scale is used the pH scale ranges from 0 to 14 a pH of 7 is neutral in an acid base reaction

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web this point of titration is referred to as the equivalence point considering the 1 1 stoichiometry of this acid base reaction $\text{NaOH} + \text{C}_6\text{H}_4\text{COOH} \rightarrow \text{NaC}_6\text{H}_4\text{COO} + \text{H}_2\text{O}$ the point of equivalence is the point of titration when the number of moles of naoh added is equal to the number of moles of khp in the

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