

## Chapter 2 Properties of Matter

**Section 2.1 Classifying Matter**

(pages 38–44)

*This section explains how materials are classified as pure substances or mixtures. It discusses types of pure substances and mixtures.*

**Reading Strategy (page 38)**

**Summarizing** As you read, complete the classification of matter in the diagram below. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

**Pure Substances (page 39)**

- Is the following sentence true or false? Every sample of a pure substance has exactly the same composition and the same properties. true
- What are the two categories of pure substances?  
a. Elements      b. Compounds

**Elements (pages 39–40)**

- What is an element? An element is a substance that cannot be broken down into simpler substances.
- Is the following sentence true or false? The smallest particle of an element is an atom. true
- Why does an element have a fixed, uniform composition? An element has a fixed composition because it contains only one type of atom.
- Circle the letter before each element that is a gas at room temperature.  
a. carbon      b. oxygen  
c. mercury      d. nitrogen

*Match each element to its correct symbol.*

Element	Symbol
<u>b.</u> 7. aluminum	a. C
<u>c.</u> 8. gold	b. Al
<u>a.</u> 9. carbon	c. Au

# Matter Physical Science Study Guide Answers

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Gardiner, Gerald Druce, H. W. Blood-  
Ryan**

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**Chemical News and Journal of Physical Science** William Crookes, James H. Gardiner, Gerald Druce, H. W. Blood-Ryan, 1916

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