

Harcourt Science Grade 1 NYC Edition Curriculum Map Units 1-4

| September - October | November-December | January-February | March-April |
|--|--|--|--|
| Unit 1 Ready, Set, Science! | Unit 2 Animal Diversity | Unit 3 Properties of Matter | Unit 4 Weather and Seasons |
| Big Idea: How Do We Use Our Senses? | Big Idea: How are animals alike and different? | Big Idea: What are some properties of solids, liquids and gases? | Big Idea: What are some of the changes we notice between seasons? |
| Essential Questions: <ul style="list-style-type: none"> How do we put our senses to daily use? How do we handle facts safely? What are inquiry skills? | Essential Questions: <ul style="list-style-type: none"> How are animals alike and different? What do animals need to survive? How do we classify animals into groups? | Essential Questions: <ul style="list-style-type: none"> What is matter? How do we describe the properties of Matter? How can we compare and contrast different kinds of matter? | Essential Questions: <ul style="list-style-type: none"> What are some of the changes we notice between seasons? How does weather affect what we wear and activities we do? |
| Standards: L1.1.a, L1.1.b, L1.1.c, L2.2.a, L2.2.b, L2.2b, L4.1.a, L4.1.g | Standards: L1.1.a, L1.1.b, L1.1.c, L2.2.a, L2.2.b, L2.2b, L4.1.a, L4.1.g | Standards: PS.1.a, PS.1.b, PS.1.c, PS.3.a, PS.3.b, PS.3.c, L4.1.g | Standards: PS.1.a, PS.1.b, PS.3.a, PS.3.b, PS.3.c, PS.4.a |
| How Do We Use Our Senses? <ul style="list-style-type: none"> Identify the five senses. Identify the body part associated with each sense. Explain how the five senses help us learn. Hide items in a bag and use your senses to figure out what the item is. How Do We Use Inquiry Skills? <ul style="list-style-type: none"> Draw conclusions based on observations. Identify inquiry skills used by scientists. Describe how to apply inquiry skills during investigations. How Do We Use Science Tools? <ul style="list-style-type: none"> List a balance to measure | What are Living and Nonliving Things? <ul style="list-style-type: none"> Classify things as living and nonliving. Explain that living things all need food, water, and air. What Do Animals Need? <ul style="list-style-type: none"> Observe animals meeting their needs. Explain the animals need food, water, air, and shelter to live. How Can We Group Animals? <ul style="list-style-type: none"> Classify animals into groups, based on observable characteristics. Identify physical characteristics of animals and their functions. | What is Matter? <ul style="list-style-type: none"> Classify solid objects based on size, shape and color. Recognize that everything around us is matter. Recognize that objects can be made up of parts too small to see without a hand lens. What Can We Observe about Solids? <ul style="list-style-type: none"> Understand that a solid is matter that keeps its shape. Measure and compare physical properties and solid properties. What Can We Observe about Liquids? <ul style="list-style-type: none"> Understand that matter flows are liquid. Understand that matter may | What is Weather? <ul style="list-style-type: none"> Know that weather is the condition of the air outside. Use a chart to observe and compare weather data and weather patterns. How Can We Measure Weather? <ul style="list-style-type: none"> Observe and measure weather by using a variety of tools. Use thermometers to measure the temperature of air. Keep a graph of the month's weather. What Makes Clouds and Rain? <ul style="list-style-type: none"> Describe how clouds form when warmer air meets cooler air. Describe how rain forms from |

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R Sanford



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kg the molality equation is a mathematical formula that describes how much a substance weighs in this equation m moles

solute per kilogramme of solution

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with 1200 ml of 2m hno₃ if the final density of solution is same as pure water then molarity of the iron in resulting solution

which decides the nature of the above solution is 1 0 25m 2 0 5m 3 0 888m 4 none of

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of the solution at that temperature m m d m m d which is more important molality or molarity both molality and molarity are

concentration units and they are used interchangeably

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web correct answer c both a and b q2 the density of 1 m solution of hcl is 1 0585 g ml the molality of the solution is 1 0585 1

0 10 0 0585 correct answer b 1 explanation molality begin array l m left frac 1000 m 1000d mm right end array

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web sep 13 2023 the only difference between molality and molarity is in the denominator molality is measured in kilograms

of solvent whereas molarity is measured in liters of solution when studying the properties of solutions related to vapour

pressure and temperature changes concentrations expressed in molality are used

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web molarity is the number of moles of a substance per litre of solution also known as molar concentration a capital M signifies solutions labelled with molar concentration a 1.0 M solution contains 1 mole of solute per litre of solution molality is the number of solvent moles per kilogram

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web 1 molarity $M = \frac{\text{mass}}{\text{volume}}$ mass $\times \frac{1}{\text{molar mass}}$ $\times \frac{1}{\text{volume}}$ $\times 1000$ g 141.9579 g mol $\times \frac{1}{0.0129 \text{ m}^3}$ molality $m = \frac{\text{mass}}{\text{mass of solvent}}$ $\times \frac{1}{\text{molar mass}}$ $\times 1000$ g 141.9579 g mol $\times \frac{1}{0.00581158 \text{ kg}}$ mol 0.00581158 mol 0.4500 kg 0.0129 m 3 mole fraction $n_2 = \frac{\text{moles of solute}}{\text{moles of solute} + \text{moles of solvent}}$ $\times 100$ hpo 4.0 825 g 141.9579 g mol 0.00581158 mol h 2.0 450.0 g 18.015 g mol 24.97918401 mol

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web apr 7 2014 molarity moles of solute liters of solution 8.4 2.2 a first convert 250 ml to liters 250 1000 = 0.25 then calculate molarity 5 moles / 0.25 liters = 20 M 3 c a solution with molarity 2 requires 2 M of NaOH per liter so 4 x 2.8 M = 11.2 M a solution of molarity 1.5 M requires 1.5 mol of NaOH to every litre of solvent

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web answer the correct option is c explanation the molarity of a mixture $M_{\text{mix}} = \frac{M_1 V_1 + M_2 V_2 + \dots + M_n V_n}{V_1 + V_2 + \dots + V_n}$ $\times 1000$ 1.344 M q_5 with increase in temperature which of these changes a molality b fraction of solute present in water

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web 1 calculate the molarity of a solution which contains 0.40 mol of a substance dissolved in 1.6 L of a solution 2 what is the molarity of a solution containing 325 g of NaCl dissolved in 750 ml of solution 3 140 g of KCl is dissolved in 600 ml of water what is the molarity 4 724.4 g of ammonium phosphate in 4500 ml of alcohol

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web determine the molarity for each of the following solutions 0.444 mol of COCl_2 in 0.654 L of solution 98.0 g of phosphoric acid H_3PO_4 in 1.00 L of solution 0.2074 g of calcium hydroxide Ca(OH)_2 in 40.00 ml of solution 10.5 kg

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