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web question 1 find the mole fraction of methanol CH_3OH and water in a solution prepared by dissolving 4.5 g of alcohol in 40 g of H_2O molar mass of H_2O is 18 g/mol and molar mass of CH_3OH is 32 g/mol solution moles of CH_3OH $\frac{4.5}{32} = 0.14$ mole moles of H_2O $\frac{40}{18} = 2.2$ moles therefore according to the equation

web we used mole fractions in gen chem 1 to describe gas mixtures where the partial pressure of a gas was its mole fraction

time the total pressure p_a $\times$ p_t mole fraction a left frac mole a moles total right sum n_1 all x_{n_1} mole fraction of substance a is often denoted by x_a and

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0 2885 its mole percent would be 28 85

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