

# Chapter 4: Supramolecular Problems - Answers

1. Assign the oxidation number to each atom in the following elements, molecules and ionic substances:

- a)  $\text{CO}_2$       b)  $\text{CH}_4$       c)  $\text{HNO}_3$       d)  $\text{Ca(OH)}_2$       e)  $\text{H}_2\text{SO}_4$       f)  $\text{Ag}$   
 $\text{C} = +4$        $\text{C} = -4$        $\text{N} = +5$        $\text{Ca} = +2$        $\text{S} = +6$        $\text{Ag} = 0$   
 $\text{O} = -2$        $\text{H} = +1$        $\text{O} = -2$        $\text{O} = -2$        $\text{O} = -2$        $\text{Ag} = 0$   
 $\text{O} = -2$
- g)  $\text{Fe}$       h)  $\text{BaCl}_2$       i)  $\text{CO}_2$       j)  $\text{H}_2\text{O}$       k)  $\text{HNO}_3$       l)  $\text{Fe}$   
 $\text{Fe} = 0$        $\text{Ba} = +2$        $\text{C} = +4$        $\text{H} = +1$        $\text{N} = +5$        $\text{Fe} = 0$   
 $\text{Cl} = -1$        $\text{O} = -2$        $\text{O} = -2$        $\text{O} = -2$        $\text{O} = -2$        $\text{Fe} = 0$   
 $\text{Cl} = -1$

2. Assign the oxidation number to each atom in the following ions:

- a)  $\text{NO}^+$       b)  $\text{CO}_3^{2-}$       c)  $\text{HNO}_2$       d)  $\text{SO}_4^{2-}$       e)  $\text{PO}_4^{3-}$       f)  $\text{HCO}_3^-$   
 $\text{N} = +2$        $\text{C} = +3$        $\text{N} = +3$        $\text{S} = +6$        $\text{P} = +5$        $\text{C} = +3$   
 $\text{O} = -2$        $\text{O} = -2$        $\text{O} = -2$        $\text{O} = -2$        $\text{O} = -2$        $\text{O} = -2$   
 $\text{O} = -2$

3. For the following reactions identify if it is an oxidation-reduction reaction. If it is a redox reaction identify the oxidizing agent, the reducing agent, the substance being oxidized and the substance being reduced. Identify the oxidation number change to help to support these identifications.



Redox rxn.  $\text{Cu}$  oxidized (oxidizing agent).  $\text{H}_2$  as  $\text{H}^+$   $\longrightarrow$   $\text{H}_2$   
 $\text{HCl}$  reduced (oxidizing agent).  $\text{H}_2$  as  $\text{H}^+$   $\longrightarrow$   $\text{H}_2$



Not a redox rxn.  $\text{Ca}$  as  $\text{Ca}^{2+}$  is both oxidized and reduced



Redox rxn.  $\text{Fe}$  oxidized (oxidizing agent).  $\text{Si}$  as  $\text{Si}^0 \longrightarrow \text{Si}^+$   
 $\text{Si}$  reduced (oxidizing agent).  $\text{Si}$  as  $\text{Si}^0 \longrightarrow \text{Si}^+$



Not a redox reaction. No change in ox. #



# Mastering Chemistry Answer Key

**RS Peters**



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