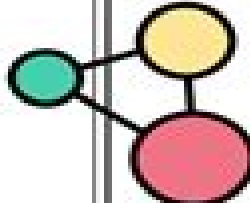
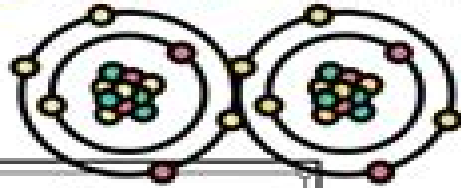


CHEMICAL BONDS

IONIC, COVALENT, & METALLIC

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READING COMPREHENSION & QUESTIONS

SCIENCE REAL

Types of Chemical Bonds Reading Comprehension

A molecule is a group of two or more atoms held together by a chemical bond. A chemical bond is a force that holds atoms together. It is sometimes thought of as a bridge or a glue that holds atoms together. There are three main types of chemical bonds: ionic, covalent, and metallic.

1. What is the difference between a molecule and a compound?

2. Which of the substances shown in the diagram above are molecules?

3. Which of the substances shown in the diagram above are compounds?

CHEMICAL BONDS AND ELECTRONEGATIVITY

Electrons in atoms exist in two main energy levels, or shells. Electrons in the outer shell are called valence electrons. The number of valence electrons in an atom determines its chemical properties. Atoms with the same number of valence electrons are in the same group on the periodic table. Atoms with different numbers of valence electrons are in different groups.

The two most common types of chemical bonds are ionic bonds and covalent bonds. Ionic bonds are formed between a metal and a nonmetal. Covalent bonds are formed between two nonmetals.

Whether a bond is ionic or covalent depends on the electronegativity of the atoms involved. Electronegativity is an atom's ability to attract electrons. Atoms with a high electronegativity have a strong pull on the electrons in a bond. Atoms with a low electronegativity have a weak pull on the electrons in a bond.

Electronegativity increases from left to right across the periodic table. It also increases from bottom to top.

1. Which atom has a higher electronegativity?

2. Which atom is a metal and which one is a nonmetal?

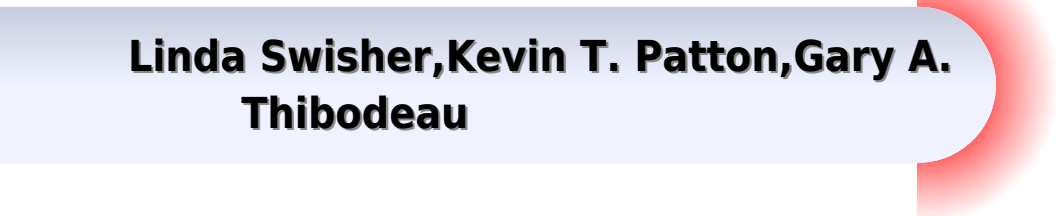
3. Which bond is ionic and which one is covalent?

4. Which bond is stronger and why?

GUIDED READING ACTIVITY

Guided Study Covalent Bonds Answer Key

**Linda Swisher, Kevin T. Patton, Gary A.
Thibodeau**



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