

Can You Find the Mole Ratio?

In this experiment, baking soda (NaHCO_3) reacts with hydrochloric acid (HCl) to produce sodium chloride, carbon dioxide, and water. You will determine the moles of reactant used and moles of product produced using both the lab data and stoichiometry and compare the two.

Data Table:

mass of test tube and baking soda	19.46 g
mass of empty test tube	19.21 g
mass of baking soda	0.25 g
mass of test tube and sodium chloride	19.39 g
mass of empty test tube	19.21 g
mass of sodium chloride	0.18 g

Conclusion Questions:

1. Calculate the number of moles of baking soda used in the lab.

$$0.25 \text{ g NaHCO}_3 \left(\frac{\text{mole}}{84.01 \text{ g}} \right) = 0.00297 \text{ moles of NaHCO}_3$$

2. Calculate the number of moles of sodium chloride produced in the lab.

$$0.18 \text{ g of NaCl} \left(\frac{\text{mole}}{58.44} \right) = 0.0031 \text{ moles of NaCl}$$

3. What is the experimental mole ratio of baking soda (NaHCO_3) to sodium chloride (NaCl)?

$$1:1$$

4. Write a balanced equation for the reaction that took place in the experiment.



5. According to the balanced equation, what is the theoretical mole ratio of baking soda to sodium chloride?

$$1:1$$

6. Was the experimental mole ratio exactly the same as the theoretical mole ratio? If not, give some possible reasons.

Yes ~ if it wasn't some reasons

side reactions

measurement

transfer

contamination

Lab Stoichiometry Datasheet Key

R Barnett

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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1. Understanding the eBook Lab Stoichiometry Datasheet Key
 - The Rise of Digital Reading Lab Stoichiometry Datasheet Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Lab Stoichiometry Datasheet Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Lab Stoichiometry Datasheet Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Lab Stoichiometry Datasheet Key
 - Personalized Recommendations
 - Lab Stoichiometry Datasheet Key User Reviews and Ratings
 - Lab Stoichiometry Datasheet Key and Bestseller Lists
5. Accessing Lab Stoichiometry Datasheet Key Free and Paid eBooks
 - Lab Stoichiometry Datasheet Key Public Domain eBooks
 - Lab Stoichiometry Datasheet Key eBook Subscription Services
 - Lab Stoichiometry Datasheet Key Budget-Friendly Options

6. Navigating Lab Stoichiometry Datasheet Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Lab Stoichiometry Datasheet Key Compatibility with Devices
 - Lab Stoichiometry Datasheet Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Lab Stoichiometry Datasheet Key
 - Highlighting and Note-Taking Lab Stoichiometry Datasheet Key
 - Interactive Elements Lab Stoichiometry Datasheet Key
8. Staying Engaged with Lab Stoichiometry Datasheet Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Lab Stoichiometry Datasheet Key
9. Balancing eBooks and Physical Books Lab Stoichiometry Datasheet Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Lab Stoichiometry Datasheet Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Lab Stoichiometry Datasheet Key
 - Setting Reading Goals Lab Stoichiometry Datasheet Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Lab Stoichiometry Datasheet Key
 - Fact-Checking eBook Content of Lab Stoichiometry Datasheet Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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