

## Experiment 1 Lab Report

### Determining the Densities of Liquids and Solids

#### OBJECTIVE:

The objective of Experiment 1 was to find the density of an unknown liquid and an unknown solid.

#### DISCUSSION:

Density is a standard physical property of a liquid or solid. The density of a subject is expressed as a ratio of the subject's mass to its' volume in the following formula:  $\text{Density} = \text{Mass} \div \text{Volume}$ . At any specified temperature and pressure the density of a sample is constant. Given the specific density of water at various temperatures and solving the density formula for volume we can accurately determine the volume of our container. The mass of the unknown samples are accurately obtained by weighing and using the measuring by difference method. After obtaining the mass and volume of the unknown sample we can calculate the density of that unknown liquid or solid. By design, Experiment 1 provides the data necessary to calculate the density of the unknown liquid and solid.

#### EXPERIMENTAL PROCEDURE:

See attached pages.

#### DATA:

The data collected and calculated that was used to determine the volume of the pycnometer used throughout the experiment is represented in Table 1:

Table 1

	Trial 1	Trial 2
Mass of empty stoppered flask	63.517g	63.515g
Mass of stoppered flask and water	122.194g	122.20g
Mass of water (calculated)	58.677g	58.692g
Temperature of water	20.9°C	21.4°C
Density of water*	0.9980 g/mL	0.9979 g/mL

\*Density of water based on temperature, retrieved from - [jupiter.plymouth.edu](http://jupiter.plymouth.edu)

The data collected and calculated that was used to determine the density of unknown liquid 'B' is represented in Table 2:

Table 2

	Trial 1	Trial 2
Mass of stoppered flask and liquid B	111.590g	111.581g
Mass of liquid (calculated)	48.073g	48.066g
Volume of liquid B (from Table 1 calculation)	58.79mL	58.82mL

# Lab Report Chemistry Matriculation Density Of Water

**George Herbert Bailey**



## Lab Report Chemistry Matriculation Density Of Water:

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*Drinking Water Chemistry* Barbara Hauser, 2018-10-03 Whether you are a new employee or seasoned professional you need easy access to the latest test methods updated quality control procedures and calculations at your fingertips You need to perform analyses quickly and easily and troubleshoot problems as they arise You need a resource that is not only informative but also practical and easy to use *Drinking Water Chemistry A Laboratory Manual* fills this need The book gives you a thorough overview of the most basic and therefore important laboratory topics such as Laboratory Safety dos and don'ts based on real experience Sampling preservation techniques online sampling and record keeping Laboratory Instruments practical use ranges principles of operation calibration conditioning useful life and replacement common quality control issues Chemical Use reagents standards indicators purpose and use chemical quality and properties avoidance of contamination molecular weight calculations Quality Control replicate analyses spiked split and reference samples percent recovery of standard standard deviation control charts and everyday quality control measures Weights and Concentrations care and analytical balances mathematical conversions among concentration units dilutions and concentration changes The remaining chapters cover test analysis including reason for the test type of sample taken treatment plant control significance expected range of results appropriate quality control procedures apparatus used reagents including function concentration and instructions for preparation procedural steps calculations and notes on possible problems and references This is a working manual meant to be kept by your side in the lab not on the shelf in an office or library You can bend it you can lay it flat you can take it anywhere you do your job Useful and practical *Drinking Water Chemistry A Laboratory Manual* provides the information you need to perform tests understand the results apply them to the determination of water quality before and after treatment and troubleshoot any problems      **Water Chemistry Laboratory Manual** David Albert Jenkins, American Association of Professors in Sanitary Engineering, 1971      **Laboratory Manual for the Examination of Water, Waste Water, and Soil** Hans Hermann Rump, 1988      *Water Chemistry Laboratory Manual* American Association of Professors in Sanitary Engineering, David Jenkins, 1972      *Water Chemistry Laboratory Manual* Association of Environmental Engineering Professors, David Jenkins, 1973      **The Chemistry of Water** Susan E. Kegley, Joy Andrews, 1998 Purity of the water supply is a pressing problem that will intensify in coming years This laboratory manual is designed as an introduction to factors

affecting water and the methods used to assess water quality Ideal for use in the laboratory portion of an introductory environmental chemistry or general chemistry course the manual combines a careful balance of wet chemistry methods with instrumental techniques such as spectroscopy and ion chromatography The Chemistry of Water can also be used in higher level analytical chemistry and instrumental methods of analysis courses

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