

ROCK CYCLE DIRECTED READING – EARTH SCIENCE

6.1 ROCKS AND THE ROCK CYCLE

1. Rock can be a collection of one or more minerals, or it might be made of
 - a. inorganic matter.
 - b. solid organic matter.
 - c. liquid organic matter.
 - d. chemicals.
2. The word igneous comes from a Latin term that means
 - a. "from fire."
 - b. "from wind."
 - c. "from rock."
 - d. "from fossils."
3. How do rocks get broken down into small fragments?
 - a. by freezing
 - b. by erosion
 - c. by deposition
 - d. by crystallization
4. The word metamorphic means
 - a. "changed from."
 - b. "to become."
 - c. "changed form."
 - d. "to form."

Match the following:

- | | |
|---------------------|---|
| 5. igneous rock | a. rock that forms when existing rock is altered |
| 6. sedimentary rock | b. molten rock |
| 7. lava | c. rock that forms when molten rock cools and hardens |
| 8. metamorphic rock | d. rock that forms when rock fragments are compressed and cemented together |
| 9. magma | e. molten rock that is exposed at Earth's surface |
| 10. sediment | f. rocks, mineral crystals, and organic matter that have been broken into fragments |

11. Define *rock cycle*. _____

12. When bits and pieces of rock are compacted or cemented, the bits and pieces become _____ rocks.

13. If sedimentary rocks are subjected to changes in temperature and pressure, the rocks may become _____ rocks.

14. Under certain temperature and pressure conditions, metamorphic rock will melt and form _____.

15. If magma cools, it turns into new _____ rock.

16. A particular body of rock does not always pass through each stage of the _____.

17. What do the physical characteristics of a rock reflect? _____

18. When rock formed under intense _____ is uplifted to Earth's surface, decreased pressure allows the joints and fractures to open.

19. Once weaknesses are exposed to air, the processes of physical and chemical _____ begin.

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This book delves into Keytrain Locating Level 6 Schoolworld An Edline Solution. Keytrain Locating Level 6 Schoolworld An Edline Solution is a crucial topic that needs to be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Keytrain Locating Level 6 Schoolworld An Edline Solution, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:
 - Chapter 1: Introduction to Keytrain Locating Level 6 Schoolworld An Edline Solution
 - Chapter 2: Essential Elements of Keytrain Locating Level 6 Schoolworld An Edline Solution
 - Chapter 3: Keytrain Locating Level 6 Schoolworld An Edline Solution in Everyday Life
 - Chapter 4: Keytrain Locating Level 6 Schoolworld An Edline Solution in Specific Contexts
 - Chapter 5: Conclusion
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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper s mass into account add these to the beaker 3 measure 50 0 ml of water in the cylinder and add it to the beaker 4

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web description great lab for use during stoichiometry unit in chemistry reinforces ideas of mole ratios and stoichiometry calculations by measuring the mass of iron that reacts and the mass of copper metal produced students can calculate the ratio of moles of reactant to moles of product this experimental mole ratio from lab can be compared

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web in this experiment we will use stoichiometric principles to deduce the appropriate equation for the reaction between metallic iron and a solution of copper ii sulfate this reaction produces metallic copper which is seen precipitating as a finely divided red powder

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web in this lab we will use experimental evidence and the principles of stoichiometry to determine which ion is formed in a single replacement reaction between iron and copper ii sulfate pre lab assignment in preparation for the lab activity we will investigate another single replacement reaction between copper and silver nitrate

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the copper ii is then reduced and gains electrons

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web stoichiometry can be defined as the relationship between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate

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web under the finalize of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away cooper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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