

Experiment: Charles's Law

Introduction:

According to Charles' Law the volume of a gas is directly proportional to the Kelvin temperature assuming the pressure is constant or

$$V = kT$$

Where, V = volume, T = the Kelvin temperature and k = a constant number

If we have two different temperatures, T_1 and T_2 Charles Law can be restated as:

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

Using Charles Law you will predict the volume the gas should be at the lower temperature based on its volume at the higher temperature. Then you will compare the predicted volume at the lower temperature with the actual measured volume.

To summarize, a given volume of air was taken at the temperature of boiling water. The air was cooled, causing it to contract. The new volume was determined experimentally.

Objective:

The purpose of this experiment is to study the changes in the volume of a gas with changes in temperature at constant pressure.

Apparatus:

Erlenmeyer Flask, 250 mL, Thermometer, Beaker, 600 mL, Graduated cylinder 100 mL, Rubber stopper fitted with a pinch clamp, Large dish pan, Ring stand and clamp, Hot plate.

Procedure (labelled flow diagram):

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S Ashworth



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A Laboratory Manual to Accompany Godfrey's Elementary Chemistry Hollis Godfrey, 1910 **General Chemistry Laboratory Manual** John Charles Blake, 1913 *A Laboratory Manual of General Chemistry* James Henri Walton, Francis Craig Krauskopf, 1920 **The Essential Lab Manual** Karen Timberlake, 2002-06-24 Drawing from the successful main Laboratory Manual the Essential Laboratory Manual includes twenty one experiments which have been revised and updated Suitable for a one or two term lab course **Laboratory Experiments to Accompany General, Organic and Biological Chemistry** Charles Anderson, David B. Macaulay, Molly M. Bloomfield, Joseph M. Bauer, 2013-02-04 This General Organic and Biochemistry text has been written for students preparing for careers in health related fields such as nursing dental hygiene nutrition medical technology and occupational therapy It is also suited for students majoring in other fields where it is important to have an understanding of the basics of chemistry An integrated approach is employed in which related general chemistry organic chemistry and biochemistry topics are presented in adjacent chapters This approach helps students see the strong connections that exist between these three branches of chemistry and allows instructors to discuss these interrelationships while the material is still fresh in students minds *Laboratory Manual for Principles of General Chemistry* Jo Allan Beran, 2010-11-01 This new edition of the Beran lab manual emphasizes chemical principles as well as techniques The manual helps students understand the timing and situations for the various techniques The Beran lab manual has long been a market leading lab manual for general chemistry Each experiment is presented with concise objectives a comprehensive list of techniques and detailed lab intros and step by step procedures *A Laboratory manual of elemental qualitative chemical analysis* Andrew Richard Bliss, 1918 **A Laboratory Manual of Physics and Applied Electricity: Senior courses and outlines of advanced work** Edward Leamington Nichols, 1896 **Laboratory Manual and Principles of Chemistry for Beginners** George Mann Richardson, 1894 *A Laboratory Manual of Qualitative Chemical Analysis* Hugo Herman Schaefer, Andrew Richard Bliss, 1928 **Laboratory Manual of Inorganic Chemistry for Colleges** Lyman Churchill Newell, 1916 *A Laboratory Manual of General Chemistry* Leon Burr Richardson, Andrew Jackson Scarlett, 1929 **Laboratory Manual for Principles of General Chemistry** J. A. Beran, Mark Lassiter, 2022-08-16 Laboratory Manual for Principles of General Chemistry 11th Edition covers two semesters of a general chemistry laboratory program The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure selecting the proper apparatus employing the proper techniques systematically analyzing and interpreting the data and minimizing inherent variables As a result of good data a scientific and analytical conclusion is made which may or may not be right but is certainly consistent with the data Experiments write textbooks textbooks don't write experiments A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered Further experimentation provides additional cause effect

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