Lab: Solubility—Datasheet

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Procedure:

1) Each lab group was assigned a different temperature, calculating the solubility of KCl in water at that temperature. Each group conducted the lab as seen on the video, and their data follows. You are to complete columns #5 and #6, using the provided data from each lab group.

Data:

#1 Assigned Temperature °C	#2 Mass of evaporating dish + cover	#3 Mass of evaporating dish + cover + KCl solution	#4 Mass of evaporating dish + cover + dry KCl	#5 Mass of KCl (#4 - #2)	#6 Mass of water (#3 - #4)	#7 Mass of KCl per 100 g of water
10	37.81 g	48.75 g	40.54 g	2.73 g	8.21g	33.25g
20	37.65 g	50.42 g	40.99 g	3.34g	9.43g	35.42g
30	36.95 g	47.82 g	39.81 g	2.86g	8.01g	35.71g
40	37.80 g	48.19 g	40.63 g	2.83g	7.56g	37.43g
50	36.50 g	47.49 g	39.75 g	3.25g	7.74g	41.99g
60	37.75 g	50.08 g	41.52 g	3.77g	8.56g	44.04g
70	35.98 g	48.29 g	39.86 g	3.88g	8.43g	46.03g
80	37.81 g	48.64 g	41.37 g	3.56g	7.27g	48.97g
90	36.99 g	48.96 g	40.98 g	3.99g	7.98g	50.0g

2) Column #7 is to be calculated as to express the solubility of KCl in grams per 100 grams of water. Following the equation given, complete column #7 for each group's data.

$$? \text{ g KCl} = 100 \text{ g H}_2\text{O} \times \frac{\text{Mass of KCl}}{\text{Mass of water}} =$$

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