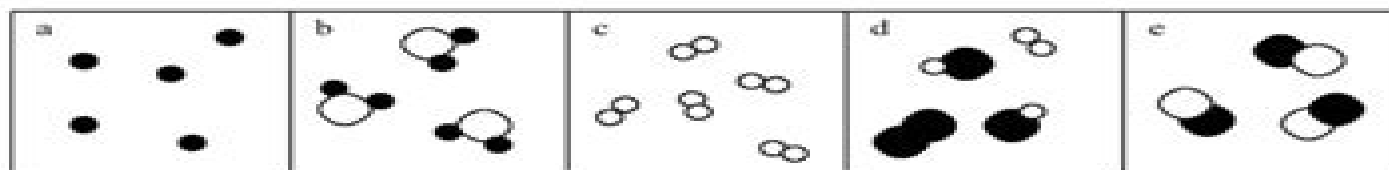


## PRACTICE TEST for HONORS CHEMISTRY EXAM 2: ATOMS

1. Prepare a summary of Chapter 4 in your text, in the format of your choice (outline, concept map . . .).



2. Describe what each box represents (e.g., drawing (a) represents a pure monatomic element)

3. Draw circle pictures to illustrate each of these.

- a. 4 molecules of  $A_2$
- b. 4 molecules of  $A_2B$
- c. 2 molecules of  $A_2$  and 2 molecules of  $B_2$
- d. 2 molecules of  $A_2B$  and 2 molecules of  $AB$

4. Fill in the blanks to complete these paragraphs. In your response, write the entire paragraph, not just the missing words.

The Greek philosopher \_\_\_\_\_ was among the first to suggest the existence of atoms, but the modern idea of the atom began with \_\_\_\_\_'s atomic theory in the year 1808. A few years later, in \_\_\_\_\_, Amadeo Avogadro stated his hypothesis that \_\_\_\_\_ to explain why gases combine in simple, whole number ratios. His hypothesis later became the basis for the mole, the counting unit of chemistry. One mole contains \_\_\_\_\_ particles.

Almost a century later, in \_\_\_\_\_, the English physicist \_\_\_\_\_ discovered the electron and showed that the atom was made of even smaller particles. He thought the atom was like a blueberry muffin, but in the year \_\_\_\_\_, \_\_\_\_\_ did an experiment in which he directed alpha particles at a thin sheet of gold foil. He expected that \_\_\_\_\_, but to his astonishment \_\_\_\_\_! From this experiment, he proposed the nuclear model of the atom, in which \_\_\_\_\_.

5.
  - a. List the 3 subatomic particles we studied, including their charge and mass in amu.
  - b. Describe the arrangement of these particles in the atom.
  - c. Compare the size of the nucleus to the size of the atom [hint: it involves football]
6. What are isotopes? How will the symbols of isotopes be alike? how will they differ?
7. Fill in this table:

Symbol	Protons	Neutrons	Electrons	Atomic number	Mass number
$^{131}_{51}\text{I}$					
	47	61			
				38	90
	9				19

8. Rubidium is 72.17%  $^{85}\text{Rb}$  (84.9117 amu) and 27.83%  $^{87}\text{Rb}$  (86.9178 amu). Calculate the average atomic mass of Rb. Don't forget units and watch your sigfigs.

# Honors Chemistry Practice Test

**LP Steffe**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the LP Steffe text.

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