

## Honors Chemistry Rate Law Homework

- 1) For the reaction  $A + B \rightarrow C$ , we've determined the following experimental data:

Time	$[A]$ (M)	$[B]$ (M)	Initial rate (M/s)
1	0.0100	0.0100	$1.82 \times 10^{-3}$
2	0.0200	0.0100	$3.24 \times 10^{-3}$
3	0.0100	0.0200	$4.48 \times 10^{-3}$

Using this information, determine the rate law for this reaction, find the reaction order for A and B, and find the overall reaction order.

- 2) In the reaction  $A + B \rightarrow C$ , we found the following rate data:

Time	$[A]$ (M)	$[B]$ (M)	Initial rate (M/s)
1	0.005	0.015	$2.85 \times 10^{-3}$
2	0.005	0.030	$5.69 \times 10^{-3}$
3	0.010	0.015	$11.3 \times 10^{-3}$

Using this information, find the overall rate law for this reaction, find the order of the reaction for each reactant, and the overall reaction order.

- 3) Using the information from the rate equations you found in problem 2, determine the rate constant  $k$  for this reaction.

# Honors Chemistry Rate Law Homework

**Lingsheng Yao**



## **Honors Chemistry Rate Law Homework:**

## Reviewing **Honors Chemistry Rate Law Homework**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is really astonishing. Within the pages of "**Honors Chemistry Rate Law Homework**," an enthralling opus penned by a highly acclaimed wordsmith, readers attempt an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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