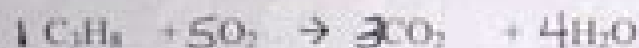


## Limiting Reagent Worksheet #1

1. Given the following reaction: (Balance the equation first!)



a) If you start with 14.8 g of  $\text{C}_3\text{H}_8$  and 3.44 g of  $\text{O}_2$ , how many grams of  $\text{H}_2\text{O}$  are produced?

$$\frac{14.8 \text{ g C}_3\text{H}_8}{44.097 \text{ g C}_3\text{H}_8} \times \frac{1 \text{ mol C}_3\text{H}_8}{1 \text{ mol C}_3\text{H}_8} \times \frac{4 \text{ mol H}_2\text{O}}{1 \text{ mol C}_3\text{H}_8} \times \frac{18.015 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = \frac{1066.488}{44.097} = 24.19 \text{ g}$$

$$\frac{3.44 \text{ g O}_2}{31.998 \text{ g O}_2} \times \frac{1 \text{ mol O}_2}{1 \text{ mol O}_2} \times \frac{4 \text{ mol H}_2\text{O}}{5 \text{ mol O}_2} \times \frac{18.015 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = \frac{247.8564}{159.99} = 1.549 \text{ g}$$

b) Which is the limiting reagent (which reactant created less  $\text{H}_2\text{O}$ )? \_\_\_\_\_

How many grams of the non-limiting reagent are actually used? \_\_\_\_\_

How many grams of the non-limiting reagent are in excess (left over)? \_\_\_\_\_

Following equation:



# Limiting Reactants Study Guide Answer Key

**N Noddings**



## Limiting Reactants Study Guide Answer Key:

*Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual* Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter. **Chemistry, Student Study Guide** James E. Brady, Fred

Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst. Carbon nanotubes are a recently discovered allotrope of carbon. Three other allotropes of carbon, buckyballs, graphite, and diamond, are illustrated at the left, as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made. The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet. The study of chemistry also follows from the simple to the more complex, and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement. [Class 8-12 Chemistry Questions and Answers PDF](#) Arshad Iqbal, The Class 8-12 Chemistry Quiz Questions and Answers PDF, Grade 8-12 Chemistry Competitive Exam Questions, Chapter 1-15 Practice Tests, Chemistry Textbook Questions for Beginners includes Questions to solve problems with hundreds of class questions. Class 8-12 Chemistry Questions and Answers PDF book covers basic concepts and analytical assessment tests. Class 8-12 Chemistry Quiz PDF book helps to practice test questions from exam prep notes. The Grade 8-12 Chemistry Quiz Questions and Answers PDF eBook includes Practice material with verbal, quantitative, and analytical past papers questions. Class 8-12 Chemistry Questions and Answers PDF Free download, chapter 1, a book to review textbook questions on chapters: Molecular structure, acids and bases, atomic structure, bonding, chemical equations, descriptive chemistry, equilibrium systems, gases, laboratory, liquids and solids, mole concept, oxidation-reduction, rates of reactions, solutions, thermochemistry. Questions for high school and college revision questions, Chemistry Interview Questions and Answers PDF Download free eBook's sample covers: beginner's solved questions, textbook's study notes to practice online tests. The Grade 8-12 Chemistry Interview Questions, Chapter 1-15 PDF book includes high school workbook

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