

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

## CHAPTER 12

## STUDY GUIDE FOR CONTENT MASTERY

## Section 12.3 Limiting Reactants

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then complete the table. Show your calculations for questions 25–27 in the space below the table.



The molar masses of  $O_2$ ,  $NO$ , and  $NO_2$  are 32.00 g/mol, 30.01 g/mol, and 46.01 g/mol, respectively.

Amount of $O_2$	Amount of $NO$	Amount of $NO_2$	Limiting Reactant	Amount and State of Excess Reactant
1 molecule	2 molecules	2 molecules	none	none
4 molecules	4 molecules	4 molecules	NO	2 molecules $O_2$
2 molecules	8 molecules	4 molecules	2. $O_2$	3. 4 molecules NO
1.00 mol	2.00 mol	4. 2.00 mol	5. none	6. none
4.00 mol	4.00 mol	7. 4.00 mol	8. NO	9. 2.00 mol $O_2$
1.00 mol	7.00 mol	10. 7.00 mol	11. NO	12. 1.50 mol $O_2$
1.00 mol	4.00 mol	13. 2.00 mol	14. $O_2$	15. 2.00 mol NO
0.500 mol	0.200 mol	16. 0.200 mol	17. NO	18. 0.400 mol $O_2$
32.00 g	60.02 g	19. 92.02 g	20. none	21. none
16.00 g	60.02 g	22. 46.01 g	23. $O_2$	24. 58.12 g NO
16.00 g	20.00 g	25. 28.76 g	26. $O_2$	27. 1.24 g NO

balanced equation mole ratio = 2 mol  $NO$ /1 mol  $O_2$

$16.00 \text{ g } O_2 \times 1 \text{ mol } O_2 / 32.00 \text{ g } O_2 = 0.5125 \text{ mol } O_2$

$20.00 \text{ g } NO \times 1 \text{ mol } NO / 30.01 \text{ g } NO = 0.6664 \text{ mol } NO$

actual mole ratio =  $0.6664 \text{ mol } NO / 0.5125 \text{ mol } O_2 = 2.132 \text{ mol } NO / 1.000 \text{ mol } O_2$

Because the actual mole ratio of  $NO$ : $O_2$  is larger than the balanced equation mole ratio of  $NO$ : $O_2$ , there is an excess of  $NO$ ;  $O_2$  is the limiting reactant.

Mass of  $NO$  used =  $0.5125 \text{ mol } O_2 \times 2 \text{ mol } NO / 1 \text{ mol } O_2 = 0.4250 \text{ mol } NO$

$0.4250 \text{ mol } NO \times 30.01 \text{ g } NO / 1 \text{ mol } NO = 12.76 \text{ g } NO$

Mass of  $NO_2$  produced =  $0.4250 \text{ mol } NO_2 \times 46.01 \text{ g } NO_2 / 1 \text{ mol } NO_2 = 19.56 \text{ g } NO_2$

Excess  $NO = 20.00 \text{ g } NO - 12.76 \text{ g } NO = 7.24 \text{ g } NO$

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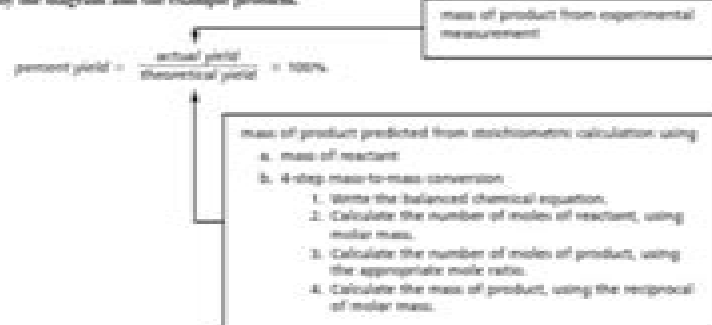
## CHAPTER 12

## STUDY GUIDE FOR CONTENT MASTERY

## Section 12.4 Percent Yield

In your textbook, read about the yield of products.

Study the diagram and the example problem.



**Example Problem:** The following chemical equation represents the production of gallium oxide, a substance used in the manufacturing of some semiconductor devices.



In one experiment, the reaction yielded 7.42 g of the oxide from a 7.00-g sample of gallium. Determine the percent yield of this reaction. The molar masses of  $Ga$  and  $Ga_2O_3$  are 69.72 g/mol and 187.44 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

1. actual yield: unknown 7.42 g  $Ga_2O_3$
2. mass of reactant: 7.00 g  $Ga$ , correct
3. number of moles of reactant:  $7.00 \text{ g } Ga \times \frac{69.72 \text{ g } Ga}{1 \text{ mol } Ga} = 7.00 \text{ g } Ga \times 1 \text{ mol } Ga / 69.72 \text{ g } Ga$
4. number of moles of product:  $0.100 \text{ mol } Ga \times \frac{2 \text{ mol } Ga_2O_3}{1 \text{ mol } Ga} = 0.100 \text{ mol } Ga \times \frac{2 \text{ mol } Ga_2O_3}{1 \text{ mol } Ga}$
5. theoretical yield:  $0.100 \text{ mol } Ga_2O_3 \times \frac{187.44 \text{ g } Ga_2O_3}{1 \text{ mol } Ga_2O_3} = 18.74 \text{ g } Ga_2O_3$ , correct
6. percent yield:  $\frac{7.42 \text{ g } Ga_2O_3}{18.74 \text{ g } Ga_2O_3} \times 100 = 7.42 \text{ g } Ga_2O_3 / 18.74 \text{ g } Ga_2O_3 \times 100$

# Limiting Reactants Study Guide

**Darrell D. Ebbing**



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**The Practice of Chemistry Study Guide & Solutions Manual** Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises CHEMISTRY HANDBOOK & STUDY GUIDE Gr11-12 NE Kevin Smith, 2024-02-01 A comprehensive summary of Grade 11 12 Physics Simple logical summaries with example exam questions and work through solutions The book covers the fundamentals of Grade 11 12 Physics and complements the material in any class text *Chemistry, Student Study Guide* James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter *Study Guide to Accompany Calculus for the Management, Life, and Social Sciences* Clyde Metz, 1984-01-01 Study Guide to Accompany Calculus for the Management Life and Social Sciences *Study Guide to Accompany Basics for Chemistry* Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that

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