

Limiting Reagents and Percentage Yield Worksheet

1. Consider the reaction



- a) 80.0 grams of iodine(V) oxide, I_2O_5 , reacts with 28.0 grams of carbon monoxide, CO . Determine the mass of iodine I_2 , which could be produced?

80 g I_2O_5	1 mol I_2O_5	1 mol I_2		XS
1	333.8 g I_2O_5	1 mol I_2O_5		
28 g CO	1 mol CO	1 mol I_2	253.8 g I_2	
1	28 g CO	5 mol CO	1 mol I_2	= 50.8 g I_2

- b) If, in the above situation, only 0.160 moles, of iodine, I_2 was produced.
i) what mass of iodine was produced?

0.160 mol I_2	253.8 g I_2	= 40.6 g I_2
	1 mol I_2	

- ii) what percentage yield of iodine was produced.

40.6 g I_2	$\times 100 =$	79.9 %
50.8 g I_2		

2. Zinc and sulphur react to form zinc sulphide according to the equation.



If 25.0 g of zinc and 30.0 g of sulphur are mixed,

- a) Which chemical is the limiting reactant?

25.0 g Zn	1 mol Zn	1 mol ZnS	= 382 mol ZnS	LR
1	65.41 g Zn	1 mol Zn		
30.0 g S	1 mol S	1 mol ZnS	= 934 mol ZnS	XS
1	32.1 g S	1 mol S		

- b) How many grams of ZnS will be formed?

382 mol ZnS	97.51 g ZnS	= 37.2 g ZnS
	1 mol ZnS	

- c) How many grams of the excess reactant will remain after the reaction is over?

$$.934 \text{ mol ZnS} - .382 \text{ mol ZnS} = .552 \text{ mol ZnS}$$

.552 mol ZnS	1 mol S	32.1 g S	= 17.7 g S
	1 mol ZnS	1 mol S	

3.

Limiting Reactant And Percent Yield Answers

**Barron's Educational Series, Mark
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