

Use the chart to determine the heat of formation for each reaction.  
Use formula  $\Delta H_{rxn} = \sum \Delta H_f \text{ products} - \sum \Delta H_f \text{ reactants}$

[illegible]

- $$2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}_{(l)}$$

$$\Delta^\circ\text{H}_{\text{rxn}} = [2\text{H}_2\text{O}] - [2\text{H}_2 + \text{O}_2]$$

$$\Delta^\circ\text{H}_{\text{rxn}} = [2(-285.8)] - [2(0) + 0]$$

$$\Delta^\circ\text{H}_{\text{rxn}} = -571.6 \text{ kJ}$$
- $$\text{CH}_4 + 2\text{O}_2 \rightarrow 2\text{H}_2\text{O}_{(g)} + \text{CO}_2$$

$$\Delta^\circ\text{H}_{\text{rxn}} = [2\text{H}_2\text{O} + \text{CO}_2] - [\text{CH}_4 + 2\text{O}_2]$$

$$\Delta^\circ\text{H}_{\text{rxn}} = [2(-285.8) + 393.5] - [-74.8 + 2(0)]$$

$$\Delta^\circ\text{H}_{\text{rxn}} = -890.3 \text{ kJ}$$
- $$\text{CuCl}_2 + 2\text{Ag}_{(s)} \rightarrow \text{Cu}_{(s)} + 2\text{AgCl} \quad (\Delta^\circ\text{H}^\circ, \text{CuCl}_2 = -215.6 \text{ kJ/mol})$$

$$\Delta^\circ\text{H}_{\text{rxn}} = [\text{Cu}_{(s)} + 2\text{AgCl}] - [\text{CuCl}_2 + 2\text{Ag}_{(s)}]$$

$$\Delta^\circ\text{H}_{\text{rxn}} = [0 + 2(-127)] - [-215.6 + 2(0)]$$

$$\Delta^\circ\text{H}_{\text{rxn}} = -38.4 \text{ kJ}$$
- $$\text{NaOH} + \text{HCl} \rightarrow \text{H}_2\text{O}_{(l)} + \text{NaCl}$$

$$\Delta^\circ\text{H}_{\text{rxn}} = [\text{H}_2\text{O}_{(l)} + \text{NaCl}] - [\text{NaOH} + \text{HCl}]$$

$$\Delta^\circ\text{H}_{\text{rxn}} = [-285.8 + -411.0] - [-426.6 + -92.3]$$

$$\Delta^\circ\text{H}_{\text{rxn}} = -177.9 \text{ kJ}$$
- $$2\text{NO} + \text{O}_2 \rightarrow 2\text{NO}_2$$

$$\Delta^\circ\text{H}_{\text{rxn}} = [2\text{NO}_2] - [2\text{NO} + \text{O}_2]$$

$$\Delta^\circ\text{H}_{\text{rxn}} = [2(33.9)] - [2(90.4) + 0]$$

$$\Delta^\circ\text{H}_{\text{rxn}} = -113.0 \text{ kJ}$$
- $$\text{N}_2\text{O} + \text{NO}_2 \rightarrow 3\text{NO} \quad (\Delta^\circ\text{H}^\circ, \text{N}_2\text{O} = 82.6 \text{ kJ/mol})$$

$$\Delta^\circ\text{H}_{\text{rxn}} = [3\text{NO}] - [\text{N}_2\text{O} + \text{NO}_2]$$

$$\Delta^\circ\text{H}_{\text{rxn}} = [3(90.4)] - [82.6 + 33.9]$$

$$\Delta^\circ\text{H}_{\text{rxn}} = 154.7 \text{ kJ}$$
- $$\text{CaC}_2 + 2\text{H}_2\text{O}_{(l)} \rightarrow \text{Ca}(\text{OH})_2 + \text{C}_2\text{H}_2 \quad (\Delta^\circ\text{H}^\circ, \text{CaC}_2 = -62.8 \text{ kJ/mol})$$

$$\Delta^\circ\text{H}_{\text{rxn}} = [\text{Ca}(\text{OH})_2 + \text{C}_2\text{H}_2] - [\text{CaC}_2 + 2\text{H}_2\text{O}_{(l)}]$$

$$\Delta^\circ\text{H}_{\text{rxn}} = [-986.6 + 226.7] - [-62.8 + 2(-285.8)]$$

$$\Delta^\circ\text{H}_{\text{rxn}} = -125.5 \text{ kJ}$$

# Heat Of Formation Worksheet With Answers

**JG Myers**



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