



Free-Response

Question 1. Properties of a Mixed Chloride

Answer the following questions about the compound.

Compound 1 is a mixed chloride.

- The only elements present in compound 1 are chlorine and sodium. The compound is a solid at room temperature and has a melting point of 1000°C. The compound is soluble in water and forms a colorless solution. The compound is a strong electrolyte in water.
- The compound is a solid at room temperature and has a melting point of 1000°C. The compound is soluble in water and forms a colorless solution. The compound is a strong electrolyte in water.
- The compound is a solid at room temperature and has a melting point of 1000°C. The compound is soluble in water and forms a colorless solution. The compound is a strong electrolyte in water.
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Question 2. Properties of a Mixed Chloride

Answer the following questions about the compound.

Compound 2 is a mixed chloride.

- The only elements present in compound 2 are chlorine and sodium. The compound is a solid at room temperature and has a melting point of 1000°C. The compound is soluble in water and forms a colorless solution. The compound is a strong electrolyte in water.
- The compound is a solid at room temperature and has a melting point of 1000°C. The compound is soluble in water and forms a colorless solution. The compound is a strong electrolyte in water.
- The compound is a solid at room temperature and has a melting point of 1000°C. The compound is soluble in water and forms a colorless solution. The compound is a strong electrolyte in water.
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