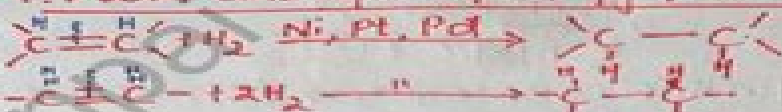


HYDROCARBON

MOP of Alkane

1. From unsaturated hydrocarbon



DM $\rightarrow$ π bond ko H.H se tod do.

exothermic

4 MCTS formed

Syn addition

Addⁿ of conj. diene



$\text{RaNi} \rightarrow$ alloy of Ni & Al in NaOH
Galvanom temp better than Ni (at 100°C)

If RaNi used $\rightarrow$ Sabatier Senderason's

$[\text{Rh}(\text{PPh}_3)_3\text{Cl}]$ used also

(Wilkinson's catalyst)

Alkene	Addition	Product
Cis	Syn	Erythro
Trans	Syn	Threo
Cis	Anti	Threo
Trans	Anti	Erythro

2. From alkyl halide



X hatoo H lagao

$\text{ROR} : \text{R} - \text{Cl} < \text{R} - \text{Br} < \text{R} - \text{I}$ $\text{ROR} \propto \frac{1}{\text{BE}}$

3. Wurtz reaction



produce higher alkane with even no. of 'C'

$\text{R} - \text{X} \rightarrow$ coupling

(1,2)

$\text{R} - \text{X} \rightarrow$ disproportionation

(3)

in 2 molecular wurtz faster than in 2 molecular but prod should be stable i.e A.A nahi banana hai.

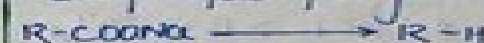
$\text{ROR} \propto \frac{1}{\text{SH}}$ $\text{R}^\cdot \text{X} > 2^\cdot \text{R}^\cdot \text{X} > 3^\cdot \text{R}^\cdot \text{X}$

for same R: $\text{RF} < \text{RCI} < \text{RBr} < \text{RI}$

4. Decarboxylation reaction



- COOH hatoo H lagao



- COONa hatoo H lagao

$\text{ROR} \propto$ stability of intermediate ($\text{R}^\cdot$)

5. Kolbe's electrolysis



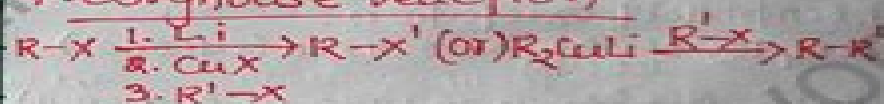
Cathode $\rightarrow \text{H}_2$ Anode $\rightarrow$ Alkane & CO_2

Addition Reaction

6. Frankland reaction



7. Coreyhouse reaction



Attack of $\text{R}^\cdot \rightarrow \text{RDS}$

$\text{ROR} \propto \frac{1}{\text{SH of R}' - \text{X}}$

If same R:

$\text{R} - \text{F} < \text{R} - \text{Cl} < \text{R} - \text{Br} < \text{R} - \text{I}$

$\text{BE} \downarrow \rightarrow$ leaving tendency $\uparrow$

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