

14.2 Math Skills

$$\text{Energy} = cm\Delta t$$

$$t_f = t_i - \frac{\text{energy}}{cm}$$

$$t_f = t_i + \frac{\text{energy}}{cm}$$

$$c = \frac{\text{energy}}{m\Delta t}$$

$$m = \frac{\text{energy}}{c\Delta t}$$

$$\Delta t = \text{---}$$

$$\Delta t = t_f - t_i$$

Specific Heat

☺ To receive any credit, you must show all your work, including givens, the correct equation, givens plugged in and answer with unit.

- Brass is an alloy made from copper and zinc. A 0.59 kg brass candlestick has an initial temperature of 98.0°C. If 2.11×10^4 J of energy is removed from the candlestick to lower its temperature to 6.8°C, what is the specific heat of brass?
- Mercury has one of the lowest specific heats. This fact added to its liquid state at most atmospheric temperatures makes it effective for use in thermometers. If 257 J of energy are added to 0.450 kg of mercury, the mercury's temperature will increase by 4.09 K. What is the specific heat of mercury?
- Bismuth's specific heat is 121 J/kg • K, the lowest of any non-radioactive metal. What is the mass of a bismuth sample if 25 J raises its temperature 5.0 K?

Math Skills Specific Heat

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Stone, Richard F. Swindell**



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