

- 7) If I have 17 liters of gas at a temperature of 67°C and a pressure of 88.89 atm, what will be the pressure of the gas if I raise the temperature to 94°C and decrease the volume to 12 liters?

$$P_1 V_1 T_2 = P_2 V_2 T_1 \quad P_2 = \frac{(88.89 \text{ atm})(17 \text{ L})(367 \text{ K})}{(12 \text{ L})(240 \text{ K})}$$

$$P_2 = \frac{P_1 V_1 T_2}{V_2 T_1}$$

$$P_2 = 192.6 \text{ atm}$$

- 18) I have an unknown volume of gas at a pressure of 0.5 atm and a temperature of 325 K. If I raise the pressure to 1.2 atm, decrease the temperature to 320 K, and measure the final volume to be 48 liters, what was the initial volume of the gas?

$$P_1 V_1 T_2 = P_2 V_2 T_1 \quad V_1 = \frac{(1.2 \text{ atm})(48 \text{ L})(325 \text{ K})}{(0.5 \text{ atm})(320 \text{ K})} = 117 \text{ L}$$

$$V_1 = \frac{P_2 V_2 T_1}{P_1 T_2}$$

- 19) The temperature inside my refrigerator is about 4°C . If I place a balloon in my fridge that initially has a temperature of 22°C and a volume of 0.5 liters, what will be the volume of the balloon when it is fully cooled by my refrigerator?

$$V_2 = ?$$

$$V_1 T_2 = V_2 T_1$$

$$\frac{V_1 T_2}{T_1} = V_2 \quad \frac{.5 \text{ L} \times 277}{295} = .47 \text{ L}$$

- 20) A man heats a balloon in the oven. If the balloon initially has a volume of 0.4 liters and a temperature of 20°C , what will the volume of the balloon be after he heats it to a temperature of 250°C ?

$$V_1 T_2 = V_2 T_1 \quad V_2 = \frac{(0.4 \text{ L})(523 \text{ K})}{(293 \text{ K})}$$

$$V_2 = \frac{V_1 T_2}{T_1}$$

$$V_2 = 0.71 \text{ L}$$

- 21) On hot days, you may have noticed that potato chip bags seem to "inflate", even though they have not been opened. If I have a 250 mL bag at a temperature of 19°C , and I leave it in my car which has a temperature of 60°C , what will the new volume of the bag be?

$$V_1 T_2 = V_2 T_1$$

$$V_2 = \frac{V_1 T_2}{T_1}$$

$$V_2 = \frac{(250 \text{ mL})(333 \text{ K})}{(292 \text{ K})} = 284.1 \text{ L}$$

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Annelies Wilder-Smith



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