

# Kinetics of Water-Rock Interaction

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# Kinetics Of Water Rock Interaction

**Clifford Lane**



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 $\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$   
 $\text{NaCl} + \text{F}_2 \rightarrow \text{NaF} + \text{Cl}_2$   
 $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$   
 $\text{Pb} + \text{HCl} \rightarrow \text{PbCl}_2 + \text{H}_2$   
 $\text{AlBr}_3 + \text{K}_2\text{SO}_4 \rightarrow \text{KBr} + \text{Al}_2\text{SO}_4$   
 $\text{CH}_4 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$   
 $\text{C}_3\text{H}_8 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

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 $\text{Na}_3\text{PO}_4 + 3\text{KOH} \rightarrow 3\text{NaOH} + \text{K}_3\text{PO}_4$   
 $\text{Mg} + \text{F}_2 \rightarrow \text{MgF}_2$   
 $\text{Li}_2\text{CO}_3 \rightarrow \text{Li}_2\text{O} + \text{CO}_2$   
 $\text{P}_2\text{O}_3 + \text{RbNO}_3 \rightarrow \text{Rb}_2\text{O} + \text{N}_2 + \text{P}_2\text{O}_5$

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 $\text{NaCl} + \text{F}_2 \rightarrow \text{NaF} + \text{Cl}_2$   
 $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$   
 $\text{Pb} + \text{HCl} \rightarrow \text{PbCl}_2 + \text{H}_2$   
 $\text{AlBr}_3 + \text{K}_2\text{SO}_4 \rightarrow \text{KBr} + \text{Al}_2\text{SO}_4$   
 $\text{CH}_4 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

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 $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$   
 $\text{HNO}_3 + \text{H}_2\text{O} \rightarrow \text{H}_2\text{O} + \text{HNO}_3$   
 $\text{H}_2\text{O} + \text{H}_2\text{O} \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}$   
 $\text{H}_2\text{O} + \text{H}_2\text{O} \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}$   
 $\text{H}_2\text{O} + \text{H}_2\text{O} \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}$

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$$\begin{array}{l} \text{N}_2 + 3 \text{H}_2 \rightarrow 2 \text{NH}_3 \\ \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2 \\ 2 \text{NaCl} + \text{F}_2 \rightarrow 2 \text{NaF} + \text{Cl}_2 \\ 2 \text{H}_2 + \text{O}_2 \rightarrow 2 \text{H}_2\text{O} \\ \text{Pb} + \text{OH}^- + 2 \text{HCl} \rightarrow 2 \text{H}_2\text{O} + 1 \text{PbCl}_2 \\ \text{AlBr}_3 + 3 \text{K}_2\text{SO}_4 \rightarrow 6 \text{KBr} + 1 \text{Al}_2\text{SO}_4 + 3 \text{CH}_4 + 2 \text{O}_2 \rightarrow 1 \text{CO}_2 + 2 \text{H}_2\text{O} \end{array}$$

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$$\begin{array}{l} 2 \text{Fe} + 3 \text{Cl}_2 \rightarrow 2 \text{FeCl}_3 \\ 4 \text{Fe} + 3 \text{O}_2 \rightarrow 2 \text{Fe}_2\text{O}_3 \\ 2 \text{FeBr}_3 + 3 \text{H}_2\text{SO}_4 \rightarrow 1 \text{Fe}_2(\text{SO}_4)_3 + 3 \text{H}_2\text{O} + 2 \text{Fe}_2\text{O}_3 + 4 \text{H}_2\text{O} \end{array}$$

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$$\begin{array}{l} 38 \text{Fe} + 2 \text{O}_3 \rightarrow 3 \text{Co} + 2 \text{Fe}_3\text{O}_4 + 3 \text{H}_2\text{O} \\ 4 \text{BaO} + \text{Ba} + 3 \text{PO}_4 + 2 \text{H}_2\text{O} + 40 \text{MgNH}_4\text{PO}_4 + 2 \text{P}_2\text{O}_7 + 2 \text{NH}_3 + \text{H}_2\text{O} + 41 \text{Al} + 3 \text{H}_2\text{SO}_4 + \text{Al}_2\text{SO}_4 + 3 \text{H}_2 + 42 \text{P}_4\text{O}_{10} + 3 \text{HCl} + \text{POCl}_3 + 3 \text{HPO}_3 + 43 \text{HNO}_2 + 2 \text{H}_2\text{O} + 3 \text{O}_2 + 4 \text{HNO}_3 \end{array}$$

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web jan 4 2019 a balanced chemical equation gives the number and type of atoms participating in a reaction the reactants products and direction of the reaction balancing an unbalanced equation is mostly a matter of making certain mass and charge are balanced on the reactants and products side of the reaction arrow

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web mar 14 2023 the key to balancing chemical equations the ultimate goal for balancing chemical equations is to make both sides of the reaction the reactants and the products equal in the number of atoms per element this stems from the universal law of the conservation of mass which states that matter can neither be created nor destroyed

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