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# Kennel Mart Online Web Page

**Bob Sehlinger**



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two substances  $x_a$  and  $x_b$  the mole fractions of each would be written as

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web solution total mass of the answer 100 g mass of benzene 30 g therefore mass of carbon tet 100 - 30 = 70 g benzene's molar mass  $\text{C}_6\text{H}_6 = 12 \times 6 + 1 \times 6 = 78 \text{ g mol}^{-1}$  hence  $\text{C}_6\text{H}_6$ 's number of moles  $30 / 78 = 0.3846 \text{ mol}$  molar mass of carbon tet  $\text{CCl}_4$  1

12 4 355 154 g mol l

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web you may work out the mole fractions on your own problem 5 what is the mass of a sample of a 0 449 molal kbr that contains 2 92 kg of water 1 molality moles solute divided by kilograms solute 0 449 mol kg x 2 92 kg x 1 31108 mol of kbr 2 moles times molar mass equals grams 1 31108 mol 119 0023 g mol 156 g kbr

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web calculate the mass percent mole fraction molality and molarity of hexane  $\text{C}_6\text{H}_{14}$  d 0 660 g  $\text{cm}^3$  when it is mixed with heptane  $\text{C}_7\text{H}_{16}$  d 0 684 g  $\text{cm}^3$  in a 35 0 55 0 ml ratio assume that the volumes are additive

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0 2885 its mole percent would be 28 85

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