

Report Table (B) B: Part C Data: pH Measurements of Dilute Solutions of Acetic Acid

Solution	pH using pH paper
0.10 M $\text{HC}_2\text{H}_3\text{O}_2$	9.0 $\pm$ 0.1 or 9.0
0.001 M $\text{HC}_2\text{H}_3\text{O}_2$	3.5 $\pm$ 0.1 or 3.5
0.0001 M $\text{HC}_2\text{H}_3\text{O}_2$	5.0 $\pm$ 0.1 or 5.0

Write the net ionic equation describing the dissociation of acetic acid in aqueous solution.

Calculate $[\text{H}^+]$ and $[\text{C}_2\text{H}_3\text{O}_2^-]$ for each concentration above, showing a sample calculation here.

$$\begin{aligned}
 \text{pH} &= -\log [\text{H}^+] & [\text{H}_2\text{PO}_4^-] &= 10^{-7} \text{ M} \\
 9.0 &= -\log x & x &= 10^{-9} = 1.0 \times 10^{-9} \text{ M} = 1.0 \times 10^{-9} \text{ M} \\
 -9.0 &= \log x & x &= 10^{-9} \\
 x &= 10^{-9} \text{ M} & K &= \frac{[1.0 \times 10^{-9} \text{ M}][1.0 \times 10^{-9} \text{ M}]}{[1.0 \times 10^{-1} \text{ M}]} \\
 x &= 1.0 \times 10^{-9} \text{ M} & K &= 1.0 \times 10^{-8} \text{ M}
 \end{aligned}$$

Write the equilibrium constant expression for acetic acid and calculate a K value and an apparent percent dissociation for each concentration, showing sample calculations here.

$$\begin{aligned}
 K_a &= \frac{[\text{H}^+][\text{C}_2\text{H}_3\text{O}_2^-]}{[\text{HC}_2\text{H}_3\text{O}_2]} = \frac{(1.0 \times 10^{-9})(1.0 \times 10^{-9})}{(1.0 \times 10^{-1})} = 1.0 \times 10^{-8} \\
 \frac{(1.0 \times 10^{-9} + 1.0 \times 10^{-9})}{1.0 \times 10^{-1}} \times 100 &= 2.0 \times 10^{-7} \%
 \end{aligned}$$

What is the effect of dilution on the percent dissociation? Why should this behavior be expected?

It In The Cards Chemistry Lab Answers

Mr. Rohit Manglik



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