

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

## CHAPTER 12

## STUDY GUIDE FOR CONTENT MASTERY

## Section 12.1 Limiting Reactants

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then complete the table. Show your calculations for questions 25–27 in the space below the table.



The molar masses of  $\text{O}_2$ ,  $\text{NO}$ , and  $\text{NO}_2$  are 32.00 g/mol, 30.01 g/mol, and 46.01 g/mol, respectively.

Amount of $\text{O}_2$	Amount of $\text{NO}$	Amount of $\text{NO}_2$	Limiting Reactant	Amount and State of Excess Reactant
1 molecule	2 molecules	2 molecules	none	none
4 molecules	4 molecules	4 molecules	NO	2 molecules $\text{O}_2$
2 molecules	8 molecules	4 molecules	2. $\text{O}_2$	3. 4 molecules NO
1.00 mol	2.00 mol	4. 2.00 mol	5. none	6. none
4.00 mol	4.00 mol	7. 4.00 mol	8. NO	9. 2.00 mol $\text{O}_2$
3.00 mol	3.00 mol	10. 3.00 mol	11. NO	12. 1.50 mol $\text{O}_2$
1.00 mol	4.00 mol	13. 2.00 mol	14. $\text{O}_2$	15. 2.00 mol NO
0.200 mol	0.200 mol	16. 0.200 mol	17. NO	18. 0.400 mol $\text{O}_2$
32.00 g	60.02 g	19. 92.02 g	20. none	21. none
16.00 g	60.02 g	22. 46.01 g	23. $\text{O}_2$	24. 50.12 g NO
16.00 g	30.00 g	25. 28.76 g	26. $\text{O}_2$	27. 1.24 g NO

balanced equation mole ratio = 2 mol  $\text{NO}$ /1 mol  $\text{O}_2$

$16.00 \text{ g } \text{O}_2 \times 1 \text{ mol } \text{O}_2/32.00 \text{ g } \text{O}_2 = 0.500 \text{ mol } \text{O}_2$

$20.00 \text{ g } \text{NO} \times 1 \text{ mol } \text{NO}/30.01 \text{ g } \text{NO} = 0.6664 \text{ mol } \text{NO}$

actual mole ratio =  $0.6664 \text{ mol } \text{NO}/0.500 \text{ mol } \text{O}_2 = 1.333 \text{ mol } \text{NO}/1.000 \text{ mol } \text{O}_2$

Because the actual mole ratio of  $\text{NO}:\text{O}_2$  is larger than the balanced equation mole ratio of  $\text{NO}:\text{O}_2$ , there is an excess of  $\text{NO}$ ;  $\text{O}_2$  is the limiting reactant.

Mass of  $\text{NO}$  used =  $0.500 \text{ mol } \text{O}_2 \times 2 \text{ mol } \text{NO}/1 \text{ mol } \text{O}_2 = 1.000 \text{ mol } \text{NO}$

$1.000 \text{ mol } \text{NO} \times 30.01 \text{ g } \text{NO}/1 \text{ mol } \text{NO} = 30.01 \text{ g } \text{NO}$

Mass of  $\text{NO}_2$  produced =  $0.500 \text{ mol } \text{O}_2 \times 46.01 \text{ g } \text{NO}_2/1 \text{ mol } \text{O}_2 = 23.01 \text{ g } \text{NO}_2$

Excess  $\text{NO} = 20.00 \text{ g } \text{NO} - 30.01 \text{ g } \text{NO} = 1.24 \text{ g } \text{NO}$

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## CHAPTER 12

## STUDY GUIDE FOR CONTENT MASTERY

## Section 12.4 Percent Yield

In your textbook, read about the yield of products.

Study the diagram and the example problem.

$$\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100\%$$

mass of product from experimental measurement

mass of product predicted from stoichiometric calculation using

- a. mass of reactant
- b. 4-step mass-to-mass conversion:
  1. Write the balanced chemical equation.
  2. Calculate the number of moles of reactant, using molar mass.
  3. Calculate the number of moles of product, using the appropriate mole ratio.
  4. Calculate the mass of product, using the reciprocal of molar mass.

**Example Problem:** The following chemical equation represents the production of gallium oxide, a substance used in the manufacturing of some semiconductor devices.



In one experiment, the reaction yielded 7.42 g of the oxide from a 3.00-g sample of gallium. Determine the percent yield of this reaction. The molar masses of Ga and  $\text{Ga}_2\text{O}_3$  are 69.72 g/mol and 187.44 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

1. actual yield: unknown  $7.42 \text{ g } \text{Ga}_2\text{O}_3$
2. mass of reactant: 3.00 g Ga correct
3. number of moles of reactant:  $3.00 \text{ g } \text{Ga} \times \frac{69.72 \text{ g } \text{Ga}}{1 \text{ mol } \text{Ga}} = 0.0430 \text{ mol } \text{Ga}$   $7.00 \text{ g } \text{Ga} \times 1 \text{ mol } \text{Ga}/69.72 \text{ g } \text{Ga}$
4. number of moles of product:  $0.0430 \text{ mol } \text{Ga} \times \frac{2 \text{ mol } \text{Ga}_2\text{O}_3}{4 \text{ mol } \text{Ga}} = 0.0215 \text{ mol } \text{Ga}_2\text{O}_3$   $0.100 \text{ mol } \text{Ga} \times \frac{2 \text{ mol } \text{Ga}_2\text{O}_3}{4 \text{ mol } \text{Ga}}$
5. theoretical yield:  $0.0215 \text{ mol } \text{Ga}_2\text{O}_3 \times \frac{187.44 \text{ g } \text{Ga}_2\text{O}_3}{1 \text{ mol } \text{Ga}_2\text{O}_3} = 4.03 \text{ g } \text{Ga}_2\text{O}_3$  correct
6. percent yield:  $\frac{7.42 \text{ g } \text{Ga}_2\text{O}_3}{4.03 \text{ g } \text{Ga}_2\text{O}_3} \times 100 = 18.4\%$   $7.42 \text{ g } \text{Ga}_2\text{O}_3/9.37 \text{ g } \text{Ga}_2\text{O}_3 \times 100$

# Limiting Reactants Study Guide

**The Princeton Review**



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