

## Limiting Reagents and Percentage Yield Worksheet

### 1. Consider the reaction



- a) 80.0 grams of iodine(V) oxide,  $\text{I}_2\text{O}_5$ , reacts with 28.0 grams of carbon monoxide,  $\text{CO}$ . Determine the mass of iodine  $\text{I}_2$ , which could be produced?

80 g $\text{I}_2\text{O}_5$	1 mol $\text{I}_2\text{O}_5$	1 mol $\text{I}_2$		XS
1	333.8 g $\text{I}_2\text{O}_5$	1 mol $\text{I}_2\text{O}_5$		
28 g $\text{CO}$	1 mol $\text{CO}$	1 mol $\text{I}_2$	253.8 g $\text{I}_2$	
1	28 g $\text{CO}$	5 mol $\text{CO}$	1 mol $\text{I}_2$	= 50.8 g $\text{I}_2$

- b) If, in the above situation, only 0.160 moles, of iodine,  $\text{I}_2$  was produced.  
i) what mass of iodine was produced?

0.160 mol $\text{I}_2$	253.8 g $\text{I}_2$	= 40.6 g $\text{I}_2$
	1 mol $\text{I}_2$	

- ii) what percentage yield of iodine was produced.

40.6 g $\text{I}_2$	$\times 100 =$	79.9 %
50.8 g $\text{I}_2$		

### 2. Zinc and sulphur react to form zinc sulphide according to the equation.



If 25.0 g of zinc and 30.0 g of sulphur are mixed,

- a) Which chemical is the limiting reactant?

25.0 g Zn	1 mol Zn	1 mol ZnS	= 382 mol ZnS	LR
1	65.41 g Zn	1 mol Zn		
30.0 g S	1 mol S	1 mol ZnS	= 934 mol ZnS	XS
1	32.1 g S	1 mol S		

- b) How many grams of  $\text{ZnS}$  will be formed?

382 mol ZnS	97.51 g ZnS	= 37.2 g ZnS
	1 mol ZnS	

- c) How many grams of the excess reactant will remain after the reaction is over?

$$.934 \text{ mol ZnS} - .382 \text{ mol ZnS} = .552 \text{ mol ZnS}$$

.552 mol ZnS	1 mol S	32.1 g S	= 17.7 g S
	1 mol ZnS	1 mol S	

### 3.

# Limiting Reactant And Percent Yield Answers

**Karin Nielsen-Saines**



## **Limiting Reactant And Percent Yield Answers:**

## Unveiling the Power of Verbal Artistry: An Mental Sojourn through **Limiting Reactant And Percent Yield Answers**

In a global inundated with displays and the cacophony of quick communication, the profound power and mental resonance of verbal artistry frequently disappear into obscurity, eclipsed by the regular onslaught of noise and distractions. However, nestled within the musical pages of **Limiting Reactant And Percent Yield Answers**, a interesting perform of fictional brilliance that impulses with organic feelings, lies an unique trip waiting to be embarked upon. Penned by way of a virtuoso wordsmith, this enchanting opus instructions viewers on a mental odyssey, lightly exposing the latent possible and profound affect stuck within the intricate web of language. Within the heart-wrenching expanse of this evocative analysis, we can embark upon an introspective exploration of the book is key subjects, dissect their charming publishing style, and immerse ourselves in the indelible impact it leaves upon the depths of readers souls.

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