

IV THERAPY TEST EXAM 1 QUESTIONS AND ANSWERS

The 5 purposes of IV therapy?

- 1) Restore or maintain _____ and _____
 - 2) Use it to _____
 - 3) Provide _____
 - 4) _____ therapy
 - 5) Route for _____ - Answers --fluid and electrolytes
- administer medicines
 - nutrition
 - Transfusion
 - blood sampling

Components of IV therapy order:

- 1) Type of _____
 - 2) Any _____ in solution and what _____ - it is (i.e. D5 1/2 NS + 20k)
 - 3) _____ and _____ and _____ (i.e. 1000mL q8h, 100mL/h)
 - 5) _____ (i.e. Peripheral IV, Central Line) - Answers --solution
- additives
 - concentration
 - Rate, volume, duration
 - Method

- Medications are administered _____ for their immediate effect.
- It is the most _____ site. (no getting those meds back)
- The absorption is most _____ (straight into blood stream)
- It is the _____, it goes right into the blood stream. You cannot get it back.
- Have to be at least _____ to do it, so if you're going straight IV push you have to be an RN, or a student with an instructor standing there - Answers --IV
- dangerous
- complete
- fastest
- RN

- We can add _____ to an infusion.
- When it comes to LPNs, if the medications are added to the infusion by the pharmacist then the LPN can _____ the IV drug. However, they are not allowed to do the actual _____.

Iv Therapy Test Questions

N Colangelo



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Table of Contents Iv Therapy Test Questions

1. Understanding the eBook Iv Therapy Test Questions
 - The Rise of Digital Reading Iv Therapy Test Questions
 - Advantages of eBooks Over Traditional Books
2. Identifying Iv Therapy Test Questions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Iv Therapy Test Questions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Iv Therapy Test Questions
 - Personalized Recommendations
 - Iv Therapy Test Questions User Reviews and Ratings
 - Iv Therapy Test Questions and Bestseller Lists
5. Accessing Iv Therapy Test Questions Free and Paid eBooks
 - Iv Therapy Test Questions Public Domain eBooks
 - Iv Therapy Test Questions eBook Subscription Services
 - Iv Therapy Test Questions Budget-Friendly Options

6. Navigating Iv Therapy Test Questions eBook Formats
 - ePub, PDF, MOBI, and More
 - Iv Therapy Test Questions Compatibility with Devices
 - Iv Therapy Test Questions Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Iv Therapy Test Questions
 - Highlighting and Note-Taking Iv Therapy Test Questions
 - Interactive Elements Iv Therapy Test Questions
8. Staying Engaged with Iv Therapy Test Questions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Iv Therapy Test Questions
9. Balancing eBooks and Physical Books Iv Therapy Test Questions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Iv Therapy Test Questions
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Iv Therapy Test Questions
 - Setting Reading Goals Iv Therapy Test Questions
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Iv Therapy Test Questions
 - Fact-Checking eBook Content of Iv Therapy Test Questions
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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$$2\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Cu(NO}_3)_2$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} = 8.3 \text{ mole } \text{H}_2\text{O}$
 $8.3 \text{ mole } \text{H}_2\text{O} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$ is your theoretical yield so percent yield $\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$
 problems 6 & 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH_3 can be produced theoretically from the reaction of 5.0 g of N what is the limiting reagent if 8.52 g are actually formed what is the

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$
 b determine the theoretical yield of Al_2O_3 $\frac{160.0 \text{ g } \text{Al}}{27 \text{ g/mol}} = 5.93 \text{ mol } \text{Al}$
 $5.93 \text{ mol } \text{Al} \times \frac{2 \text{ mol } \text{Al}_2\text{O}_3}{4 \text{ mol } \text{Al}} = 2.96 \text{ mol } \text{Al}_2\text{O}_3$
 $2.96 \text{ mol } \text{Al}_2\text{O}_3 \times 101.96 \text{ g/mol} = 301.8 \text{ g } \text{Al}_2\text{O}_3$
 c determine the percent yield $\frac{260.0 \text{ g}}{301.8 \text{ g}} \times 100 = 86.01\%$

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web 7 Na + 2 C_2O_4 is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4 33 g of H_2 are left over 13

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