- Saponifiable
- Saturated solution
- Solid
- Solubility
- Solute
- Solution
- Solvent
- Specific heat

- Standard temperature and pressure
- Stoichiometry
- Sublimation
- Substrate
- · Supersaturated solution
- Surface tension
- Temperature
- Theoretical yield

- Thermochemical equation
- Thermochemistry
- Titration
- Triprotic acid
- Unsaturated solution
- Vapor pressure lowering
- Vaporization
- Viscosity

CONCEPT QUESTIONS: You should be able to answer/describe each of the following on order to gain a real understanding of the different concepts we have covered this second semester. You can expect to see a majority of the questions on the exam similar to the ones below.

Unit 5 (Chapters 6 & 7)

- Convert each of the following to either moles or particles.
 - a. 6.022 x 10²⁹ atoms He
 - 6.022 XIO23 atoms He x I mol He
 - 2.56 mol Ne x 6.022 × 1023 atoms Ne =
- 2. Find the molar mass for each of the following compounds.
 - a. CHCl₃
 - 12.01 + 1.01+ 35.45 (3)=119.37 g/mol
 - 1.01 + 54.94 + 16.00 (4) = 1119.95 g/mol
- 3. Convert each of the following to either moles or mass.
 - 16.35g PbOx 1mol PbO =
 - -025 mol Na2504 x 142.05 a Na2504 = 13.55 g Na250
- 4. Convert each of the following to moles or volume.
 - 22.4mol H2 x 22.4L H2 = 501.76 L H2
 - . 145L F2 x 1mol F2 = 1.01 mol F2

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