

Name: _____

Data

1. mass of clean empty crucible after heating 15.4709g
2. mass of crucible with magnesium strip 15.5854g
3. mass of magnesium strip 0.1145g
4. mass of crucible and product (magnesium oxide) 15.6589g
5. mass of product 0.0735

Calculations

Label the type of chemical reaction, balance the chemical equation and calculate the Theoretical Yield of Magnesium Oxide based upon your amount of starting material.

Balance the Equation: $2\text{Mg (s)} + \text{O}_2 \text{(g)} \rightarrow 2\text{MgO (s)}$

Reaction Type: Redox

Show all calculations for the theoretical yield of magnesium oxide

$$\begin{aligned} \text{molar mass} &= \frac{80\text{g}}{2 \times 24 + 48} = \frac{25}{60} = .3125 \\ \frac{0.0735}{.3125} &= .2352 \end{aligned}$$

Theoretical Yield of Magnesium Oxide = _____

Actual Yield of Magnesium Oxide = _____

Calculation of the % Yield:

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Mark A. Griep



Gpc Chem 11511 Lab Manual:

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Table of Contents Gpc Chem 1151l Lab Manual

1. Understanding the eBook Gpc Chem 1151l Lab Manual
 - The Rise of Digital Reading Gpc Chem 1151l Lab Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Gpc Chem 1151l Lab Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gpc Chem 1151l Lab Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gpc Chem 1151l Lab Manual

- Personalized Recommendations
- Gpc Chem 1151l Lab Manual User Reviews and Ratings
- Gpc Chem 1151l Lab Manual and Bestseller Lists
- 5. Accessing Gpc Chem 1151l Lab Manual Free and Paid eBooks
 - Gpc Chem 1151l Lab Manual Public Domain eBooks
 - Gpc Chem 1151l Lab Manual eBook Subscription Services
 - Gpc Chem 1151l Lab Manual Budget-Friendly Options
- 6. Navigating Gpc Chem 1151l Lab Manual eBook Formats
 - ePub, PDF, MOBI, and More
 - Gpc Chem 1151l Lab Manual Compatibility with Devices
 - Gpc Chem 1151l Lab Manual Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gpc Chem 1151l Lab Manual
 - Highlighting and Note-Taking Gpc Chem 1151l Lab Manual
 - Interactive Elements Gpc Chem 1151l Lab Manual
- 8. Staying Engaged with Gpc Chem 1151l Lab Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gpc Chem 1151l Lab Manual
- 9. Balancing eBooks and Physical Books Gpc Chem 1151l Lab Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gpc Chem 1151l Lab Manual
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Gpc Chem 1151l Lab Manual
 - Setting Reading Goals Gpc Chem 1151l Lab Manual
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Gpc Chem 1151l Lab Manual

- Fact-Checking eBook Content of Gpc Chem 1151l Lab Manual
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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