

# Acids and bases

## Chapter outline

- Recall the definitions of Brønsted–Lowry acids and bases and deduce these in a chemical reaction, including the identification of conjugate acid–base pairs.
- Recall the terms ‘amphoteric’ and ‘amphiprotic’.
- Describe the reactions of acids with metals, metal oxides, metal hydroxides, hydrogen carbonates and carbonates.
- Recall that pH is a measure of the concentration of aqueous hydrogen ions and solve problems involving pH and the concentration of hydrogen ions and hydroxide ions.
- Explain the difference between strong and weak acids and bases and describe experiments that could be used to distinguish these.
- Recall what is meant by acid deposition and explain some of the problems associated with acid deposition as well as pre- and post-combustion methods of reducing sulfur oxide emissions.
- Recall the definitions of Lewis acids and bases and recognise these in a chemical reaction. **HL**
- Deduce the dissociation constant for weak acids and bases. **HL**
- Sketch and explain the important features of titration curves including the composition and action of a buffer and use these curves to deduce the  $pK_a$  or  $pK_b$  value of a weak acid or base. **HL**
- Explain how an acid–base indicator works and select a suitable indicator for a titration. **HL**

## KEY TERMS

**Acid deposition:** A more general term than acid rain; it refers to any process in which acidic substances leave the atmosphere and are deposited on the Earth’s surface.

**Brønsted–Lowry acid:** A proton ( $H^+$ ) donor.

**Brønsted–Lowry base:** A proton ( $H^+$ ) acceptor.

**pH:** A measure of the concentration of hydrogen ions in an aqueous solution.  
 $pH = -\log_{10}[H^+]$

**Strong acid/base:** An acid/base that is completely dissociated into its ions in aqueous solution.

**Weak acid/base:** An acid/base that is only partially dissociated into its ions in aqueous solution.

# Guided Study Workbook Acids And Bases Answers

**RJ Shavelson**



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