

HL Paper 2

Enthalpy changes depend on the number and type of bonds broken and formed.

The table lists the standard enthalpies of formation, $\Delta H_f^\ominus$, for some of the species in the reaction above.

	$\text{CH}_4(\text{g})$	$\text{H}_2\text{O}(\text{g})$	$\text{CO}(\text{g})$	$\text{H}_2(\text{g})$
$\Delta H_f^\ominus / \text{kJ mol}^{-1}$	-74.0	-242	-111	

Enthalpy changes depend on the number and type of bonds broken and formed.

a. Hydrogen gas can be formed industrially by the reaction of natural gas with steam. [3]



Determine the enthalpy change, ΔH , for the reaction, in kJ, using section 11 of the data booklet.

Bond enthalpy for $\text{C}\equiv\text{O}$: 1077 kJ mol^{-1}

b.i. Outline why no value is listed for $\text{H}_2(\text{g})$. [1]

b.ii. Determine the value of $\Delta H^\ominus$, in kJ, for the reaction using the values in the table. [1]

c. The table lists standard entropy, $S^\ominus$, values. [1]

	$\text{CH}_4(\text{g})$	$\text{H}_2\text{O}(\text{g})$	$\text{CO}(\text{g})$	$\text{H}_2(\text{g})$
$S^\ominus / \text{J K}^{-1} \text{mol}^{-1}$	+186	+189	+108	+131

Calculate the standard entropy change for the reaction, $\Delta S^\ominus$, in J K^{-1} .



d. Calculate the standard free energy change, $\Delta G^\ominus$, in kJ, for the reaction at 298 K using your answer to (b)(ii). [1]

e. Determine the temperature, in K, above which the reaction becomes spontaneous. [1]

Markscheme

a. bonds broken: $4(\text{C}-\text{H}) + 2(\text{H}-\text{O})/4(414) + 2(463)/2582 \text{ kJ}$

bonds made: $3(\text{H}-\text{H}) + \text{C}\equiv\text{O}/3(436) + 1077/2385 \text{ kJ}$

$\Delta H = \text{IBE}_{\text{bonds broken}} - \text{IBE}_{\text{bonds made}} = 2582 - 2385 = -197 \text{ kJ}$

Award [3] for correct final answer.

Award [2 max] for -197 kJ.

[3 marks]

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Table of Contents M12 Ms Mathematic Hl Paper2 Tz1

1. Understanding the eBook M12 Ms Mathematic Hl Paper2 Tz1
 - The Rise of Digital Reading M12 Ms Mathematic Hl Paper2 Tz1
 - Advantages of eBooks Over Traditional Books
2. Identifying M12 Ms Mathematic Hl Paper2 Tz1
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an M12 Ms Mathematic Hl Paper2 Tz1
 - User-Friendly Interface
4. Exploring eBook Recommendations from M12 Ms Mathematic Hl Paper2 Tz1
 - Personalized Recommendations
 - M12 Ms Mathematic Hl Paper2 Tz1 User Reviews and Ratings
 - M12 Ms Mathematic Hl Paper2 Tz1 and Bestseller Lists
5. Accessing M12 Ms Mathematic Hl Paper2 Tz1 Free and Paid eBooks
 - M12 Ms Mathematic Hl Paper2 Tz1 Public Domain eBooks
 - M12 Ms Mathematic Hl Paper2 Tz1 eBook Subscription Services
 - M12 Ms Mathematic Hl Paper2 Tz1 Budget-Friendly Options
6. Navigating M12 Ms Mathematic Hl Paper2 Tz1 eBook Formats
 - ePub, PDF, MOBI, and More
 - M12 Ms Mathematic Hl Paper2 Tz1 Compatibility with Devices
 - M12 Ms Mathematic Hl Paper2 Tz1 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of M12 Ms Mathematic Hl Paper2 Tz1
 - Highlighting and Note-Taking M12 Ms Mathematic Hl Paper2 Tz1
 - Interactive Elements M12 Ms Mathematic Hl Paper2 Tz1
8. Staying Engaged with M12 Ms Mathematic Hl Paper2 Tz1

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers M12 Ms Mathematic Hl Paper2 Tz1
- 9. Balancing eBooks and Physical Books M12 Ms Mathematic Hl Paper2 Tz1
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection M12 Ms Mathematic Hl Paper2 Tz1
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine M12 Ms Mathematic Hl Paper2 Tz1
 - Setting Reading Goals M12 Ms Mathematic Hl Paper2 Tz1
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of M12 Ms Mathematic Hl Paper2 Tz1
 - Fact-Checking eBook Content of M12 Ms Mathematic Hl Paper2 Tz1
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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