# ACIDS, BASES AND SALTS

Definition of an acid	What causes the acidic properties of acids?	
An acid is a compound which when dissolved in water produces hydrogen ions as the only positively charged ions	The hydrogen ions (H') cause the acidic properties and these are formed in the presence of water.	

Another term to refer to an acid	Why an acid is also called a proton donor?	
An acid is called a proton donor	It's because an acid provides protons or hydrogen ions (H·) to other	
	substances during the reaction.	

Substance to which an	A base is then called	Why the base is called a	Equation for the reaction between
acid provides protons		proton acceptor?	acid and base
BASE	PROTON ACCEPTOR	Because the base accepts hydrogen ions from acids	H'mg + OHmg H <sub>2</sub> O <sub>R</sub>

Common laboratory acids		These three common laboratory acids are also called	Why these acids are also called mineral acids?	
Hydrochloric acid	HC/		They are derived from mineral salts ie chlorides for	
Sulphuric acid	H;80.	MINERAL ACIDS	HCl, sulphates for H <sub>2</sub> SO <sub>4</sub> and nitrates for HNO <sub>2</sub>	
Nitric acid	HNO <sub>5</sub>			

Other minera known		Mineral salts from which the acid is derived	Organic acids known	Naturally occurring acids
Sulphurous acid	H <sub>2</sub> SO <sub>3</sub>	Derived from sulphites	Ethanoic acid	CITRIC ACID from lemons
Carbonic acid	H <sub>2</sub> CO <sub>2</sub>	Derived from carbonates	(CH <sub>3</sub> COOH)	TARTARIC ACID from grapes
Phosphoric acid	H <sub>2</sub> PO <sub>4</sub>	Derived from phosphates		ACETIC ACID from vinegar
Nitrous acid	HNO <sub>3</sub>	Derived from nitrites	Methanoic acid	LACTIC ACID from sour milk
	Marie Control		(HCOOH)	Hydrochloric acid from digestive juices

Whenever an acid is dissolved in water, it produces	Term given to the number of hydrogen ions produced by one molecule of an acid	Definition of basicity of an acid
HYDROGEN IONS	BASICITY OF AN ACID	BASICTTY of an acid is the <u>number of</u> <u>hydrogen ions</u> produced by <u>one</u> <u>molecule</u> of an acid <u>in aqueous</u> <u>solution</u> .

# **Guide Acids Bases And Salts**

**IM Harris** 

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