

Gamma ray is charge less and massless photon that's why they have more penetrating power than alpha particle which have charge and mass and interact with matter so have less penetrating power.

18.4. What is the difference between natural and artificial radioactivity?

Ans: In natural radioactivity atomic nuclei want to become stable that's why they emit some radiations and such atoms are called radioactive and the process is called natural radioactivity.

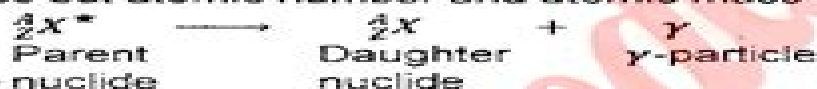
In artificial radioactivity the atomic nuclei which are not excited but we excite them by bombarding neutrons or protons and after excitation of nuclei the atom becomes radioactive and emits radiation. This is called artificial radioactivity.

18.5. How long would you likely have to wait to watch any sample of radioactive atoms completely decay?

Ans: Theoretically it will take an infinite time to decay completely because it follows the exponential law.

18.6. Which type of natural radioactivity leaves the number of protons and the number of neutrons in the nucleus unchanged?

Ans: If gamma rays emit from the excited nucleus then only energy of the nucleus is released but atomic number and atomic mass remain the same.



Example:



18.7. How much of a 1g sample of pure radioactive matter would be left after four half-lives?

Ans: Amount of sample of pure radioactive matter = 1g
 After 1st half life $1T_{1/2}$ = $\frac{1}{2}$ = 0.5g
 After 2nd half life $2T_{1/2}$ = $\frac{0.5}{2}$ = 0.25g
 After 3rd half life $3T_{1/2}$ = $\frac{0.25}{2}$ = 0.125g
 After 4th half life $4T_{1/2}$ = $\frac{0.125}{2}$ = 0.0625g

18.8. Tritium ${}^3_1\text{H}$ is a radioactive isotope of hydrogen. It decays by emitting an electron. What is the daughter nucleus?

Ans: ${}^A_Z X \longrightarrow {}^A_Z Y + {}^0_{-1}e + \text{Energy}$



18.9. What information about the structure of the nitrogen atom can be obtained from its nuclide ${}^{14}_7\text{N}$? In what way is the atom in ${}^{14}_7\text{N}$ different from the atom in ${}^{16}_7\text{N}$?

Ans: The nuclide ${}^{14}_7\text{N}$ is one of the isotopes of the nitrogen atom. It has 7 protons, 7 electrons, and 7 neutrons. Whereas ${}^{16}_7\text{N}$ isotope has two extra neutrons in its nucleus as its atomic mass number increases by 2.

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