

its second half equivalence point?

- A. $[\text{CO}_3^{2-}] = \text{maximum value}$
- B. $[\text{HCO}_3^-] = \text{maximum value}$
- C. $[\text{H}_2\text{CO}_3] = [\text{HCO}_3^-]$
- D. $[\text{HCO}_3^-] = [\text{CO}_3^{2-}]$
- E. $[\text{H}_2\text{CO}_3] = [\text{CO}_3^{2-}]$

Answer: D. $[\text{HCO}_3^-] = [\text{CO}_3^{2-}]$

4. The K_a for HCN is 6.2×10^{-10} , If there is a solution of 2M HCN, what concentration of NaCN is needed in order for the pH to be 9.2?

- A. 3M
- B. No NaCN needs to be added
- C. 2M
- D. 1M

Answer: C. 2M

5. NaOH is added to a 500ml of 2M acetic acid. If the $\text{p}K_a$ value of acetic acid is approximately 4.8, what volume of 2M NaOH must be added so that the pH of the solution is 4.8?

- A. 2L
- B. 500ml
- C. 250mL
- D. 1L

Answer: C. 250ml

6. Consider the titration curve, in which an acid is treated with aqueous 0.5M NaOH. What is the approximate $\text{p}K_a$ of the starting acid?

- A. 4.2
- B. 3.0
- C. 8.9
- D. 5.0
- E. 7.0

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Table of Contents Gre Chemistry Study Guide

1. Understanding the eBook Gre Chemistry Study Guide
 - The Rise of Digital Reading Gre Chemistry Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Gre Chemistry Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gre Chemistry Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gre Chemistry Study Guide
 - Personalized Recommendations
 - Gre Chemistry Study Guide User Reviews and Ratings
 - Gre Chemistry Study Guide and Bestseller Lists
5. Accessing Gre Chemistry Study Guide Free and Paid eBooks
 - Gre Chemistry Study Guide Public Domain eBooks
 - Gre Chemistry Study Guide eBook Subscription Services
 - Gre Chemistry Study Guide Budget-Friendly Options

6. Navigating Gre Chemistry Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Gre Chemistry Study Guide Compatibility with Devices
 - Gre Chemistry Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gre Chemistry Study Guide
 - Highlighting and Note-Taking Gre Chemistry Study Guide
 - Interactive Elements Gre Chemistry Study Guide
8. Staying Engaged with Gre Chemistry Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gre Chemistry Study Guide
9. Balancing eBooks and Physical Books Gre Chemistry Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gre Chemistry Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gre Chemistry Study Guide
 - Setting Reading Goals Gre Chemistry Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gre Chemistry Study Guide
 - Fact-Checking eBook Content of Gre Chemistry Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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