

Guided Reading Activity

The Vietnam War, 1954–1975

Lesson 2 Vietnam Divides the Nation

Review Questions: Outlining

DIRECTIONS: Read the lesson and complete the outline below. Refer to your textbook to fill in the blanks.

I. The Antiwar Movement Emerges

- A. As the Vietnam War dragged on, _____ support _____ waned because many believed that ~~the Johnson administration~~ _____ was not telling the truth about the war.
- B. Many Americans believed a _____ ~~credibility gap~~ _____ existed because of inconsistencies between _____ ~~what was shown by media~~ _____ and what the administration said about the war.
- C. Students and teachers at universities held _____ ~~teach-ins~~ _____ to explain why they opposed the war, while many protesters focused on what they thought was an unfair _____ ~~draft system~~ _____. Most of the soldiers were _____ ~~working-class and minority~~ _____ youths.
- D. As the war went on, many young men burned their _____ ~~draft cards~~ _____ while an estimated 500,000 _____ ~~refused to go~~ _____.

Many soldiers argued that if they were old enough to fight, they were old enough to

_____ ~~vote~~ _____. The _____ ~~26th amendment~~ _____ gave 18-year-olds the

- a. _____ ~~1964 Civil Rights Act and Voting Rights Act~~ _____ supported the _____ ~~war~~ _____ and those who _____ ~~opposed the war~~ _____.

B. 1968: The Peace Year

- A. Senator Eugene _____ ~~McCarthy~~ _____ argued an unconstitutional attack _____ and South Vietnam _____ This attack, called the _____ ~~Tet Offensive~~ _____, was reported. Americans, however, were alarmed that an enemy _____ ~~could launch such a large-scale attack~~ _____.

On March 25, 1968, Johnson declared the nation with the announcement that

_____ ~~he would not seek reelection~~ _____.

His announcement marked the nation _____ ~~MLK Jr.~~ _____.

- Bobby Kennedy _____ ~~who appeared to be in the lead for the~~ _____ nomination for president.

- C. _____ ~~Democratic Party~~ _____ in the presidential election of 1968. Republican candidate _____ ~~because he promised to~~ _____ the _____ ~~bring order~~ _____ to the country.

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M Mark



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web objective to apply stoichiometric principles to calculate the appropriate chemical equation for the reaction between iron
metal and a solution of copper i sulfate introduction in the form of finely split powder copper metal is produced as a
byproduct of the process as iron ions the iron metal dissolves

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web purpose to enhance the understanding of stoichiometry a reaction between iron and copper ii sulfate solution will be conducted this will help you to differentiate limiting and excess reactant in a chemical reaction finally the theoretical and percent yield of this reaction will be calculated

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web purpose in this experiment you will observe a single replacement reaction you will also use stoichiometry to predict what the theoretical yield of product is and calculate a percent yield reaction iron copper ii sulfate copper iron ii sulfate pre lab questions write the balanced equation

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web october 3 2017 abstract the techniques used in this lab are quantitative transfer and vacuum filtration with the reaction of 8 001 grams of copper ii sulfate CuSO_4 and 2 0153 grams of iron powder Fe the goal of this experiment was to determine the product of copper ii sulfate with iron

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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper s mass into account add these to the beaker 3 measure 50 0 ml of water in the cylinder and add it to the beaker 4

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web stoichiometry lab the reaction of iron with copper ii sulfate the study of stoichiometry deals with the calculation of quantities in a chemical reaction how much product will be produced how much reactant do you need to make that much product

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web description great lab for use during stoichiometry unit in chemistry reinforces ideas of mole ratios and stoichiometry calculations by measuring the mass of iron that reacts and the mass of copper metal produced students can calculate the ratio of moles of reactant to moles of product this experimental mole ratio from lab can be compared

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web in this experiment we will use stoichiometric principles to deduce the appropriate equation for the reaction between

metallic iron and a solution of copper ii sulfate this reaction produces metallic copper which is seen precipitating as a finely divided red powder

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web in this lab we will use experimental evidence and the principles of stoichiometry to determine which ion is formed in a single replacement reaction between iron and copper ii sulfate pre lab assignment in preparation for the lab activity we will investigate another single replacement reaction between copper and silver nitrate

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the copper ii is then reduced and gains electrons

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web stoichiometry can be defined as the relationship between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate

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web stoichiometric principles can be used to determine the chemical equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction the iron ions dissolve in the copper ii sulfate solution to produce copper ii ions

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web under the finalise of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away copper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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