

# Temperature, Energy, and Heat Study Guide

## I. Basic Definitions

- **Temperature:**
  - A measure of the average kinetic energy of the particles in a substance.
  - Indicates how hot or cold something is.
  - Measured in Celsius ( $^{\circ}\text{C}$ ), Fahrenheit ( $^{\circ}\text{F}$ ), or Kelvin (K).
  - SI unit is Kelvin (K).
- **Heat:**
  - The transfer of thermal energy between objects at different temperatures.
  - Flows from hotter to colder objects.
  - Measured in joules (J).
- **Thermal Energy:**
  - The total kinetic and potential energy of all the particles in a substance.
  - Depends on temperature and the amount of substance.
- **Absolute Zero:**
  - The theoretical temperature at which all molecular motion stops (0 K,  $-273.15^{\circ}\text{C}$ ).

## II. Temperature Scales

- **Celsius ( $^{\circ}\text{C}$ ):**
  - Water freezes at  $0^{\circ}\text{C}$  and boils at  $100^{\circ}\text{C}$ .
- **Fahrenheit ( $^{\circ}\text{F}$ ):**
  - Water freezes at  $32^{\circ}\text{F}$  and boils at  $212^{\circ}\text{F}$ .
- **Kelvin (K):**
  - Absolute temperature scale.
  - 0 K is absolute zero.<sup>†</sup>
  - $\text{K} = ^{\circ}\text{C} + 273.15$

## III. Heat Transfer

- **Conduction:**
  - Transfer of heat through direct contact.
  - Occurs primarily in solids.
  - Good conductors (metals) transfer heat quickly.
  - Insulators (wood, plastic) resist heat transfer.
- **Convection:**
  - Transfer of heat through the movement of fluids (liquids or gases).
  - Hotter, less dense fluids rise; cooler, denser fluids sink.
  - Creates convection currents.

# Heat Study Guide

**Michael Hoelscher**



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